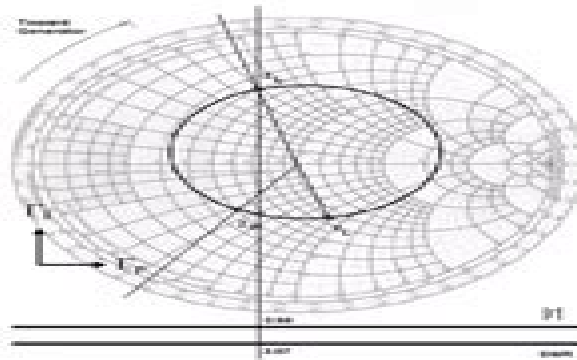


These are two families of circles as r_L and x_L . Superposition of Smith Chart and its 180° ($\lambda/4$) rotated version is known as *Combined Impedance-Admittance Smith Chart*.



$\lambda/2$ is the complete revolution of Smith chart. $\lambda/4$ is the half of Smith chart (180°). The images of Z is Y in Smith chart.

2.3. Slotted Line

This device is used to find Z_L as first V_{min} .

- Measurement of SWR on the l_{min} distance from the line.
- Calculate $|\Gamma| = (SWR - 1)/(SWR + 1)$
- $e^{j(\theta - 2\pi k l)} \big|_{V_{min}} = -1 \Rightarrow \theta = \pi + 2k l_{min}$
- Using θ and $|\Gamma|$, write $\Gamma = |\Gamma| e^{j\theta}$
- Calculate $Z_L = Z_0 [(1 + \Gamma)/(1 - \Gamma)]$ at $l = 0$.

2.4. Generator & Load Mismatches



$$v(x) = V_0^+ e^{-j\beta x} + V_0^- e^{+j\beta x} = V_0^+ (e^{-j\beta x} + \Gamma_L e^{j\beta x})$$

$$v(-l) = V_0^+ (e^{j\beta l} + \Gamma_L e^{-j\beta l}) = \frac{Z_{in}}{Z_0 + Z_g} V_g$$

$$Z_{in} = Z_0 \frac{1 + \Gamma_L e^{-j2\beta l}}{1 - \Gamma_L e^{-j2\beta l}} \bigg|_{x=-l}$$

Then, using this

$$V_0^+ = \frac{Z_0}{Z_0 + Z_g} V_g \frac{e^{-j\beta l}}{(1 - \Gamma_L \Gamma_g e^{-j2\beta l})}$$

where $\Gamma_g = Z_g - Z_0 / Z_g + Z_0$ and $\Gamma_L = Z_L - Z_0 / Z_L + Z_0$.

$$P = \frac{1}{2} \operatorname{Re}\{v_{in} i_{in}^*\} = \frac{|V_g|^2}{2} \frac{R_{in}}{(R_{in} + R_g)^2 + (X_{in} + X_g)^2}$$

where $Z_{in} = R_{in} + jX_{in}$ and $Z_g = R_g + jX_g$. Generally Z_g is fixed and three cases are considered as

- Load Matched to Line: $Z_L = Z_0$, $\Gamma_L = 0$, $SWR = 1$. Then $Z_{in} = Z_0$.

$$P = \frac{|V_g|^2}{2} \frac{Z_0}{(Z_0 + R_g)^2 + X_g^2}$$

- Generator Matched to Line: $Z_{in} = Z_g$, $\Gamma_g = 0$.

$$P = \frac{|V_g|^2}{2} \frac{R_g}{4(R_g^2 + X_g^2)}$$

- Conjugate Matching: $Z_{in} = Z_g^*$

$$\frac{\partial P}{\partial R_{in}} = 0 \Rightarrow \frac{R_{in} = R_g}{X_{in} = -X_g} \Rightarrow Z_{in} = Z_g^*$$

$$P = \frac{|V_g|^2}{2} \frac{1}{4R_g^2}$$

Maximum power transfer $\Gamma = \Gamma_g = \Gamma_L = 0$. If one directly chose $\Gamma = \Gamma_g = \Gamma_L = 0$, it does not mean that the best efficiency due to the phase differences. The efficiency can be improved only by making Z_g as small as possible.

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separately and to base its analysis on a major workshop This book presents summaries of discussions at the workshop which included sessions on the measurements and sensors originally planned for NPOESS and GOES R generation of climate data records mitigation options including the role of international partners and cross cutting issues **Global Developments in Environmental Earth Observation from Space** EARSeL Symposium,2006 *Ensuring the Climate Record from the NPOESS and GOES-R Spacecraft* National Research Council,Division on Engineering and Physical Sciences,Space Studies Board,Committee on a Strategy to Mitigate the Impact of Sensor Descope and Demanifests on the NPOESS and GOES-R Spacecraft,2008-09-25 In 2000 the nation s next generation National Polar orbiting Operational Environmental Satellite System NPOESS program anticipated purchasing six satellites for 6.5 billion with a first launch in 2008 By November 2005 however it became apparent that NPOESS would overrun its cost estimates by at least 25 percent In June 2006 the planned acquisition of six spacecraft was reduced to four the launch of the first spacecraft was delayed until 2013 and several sensors were canceled or descope in capability Based on information gathered at a June 2007 workshop Options to Ensure the Climate Record from the NPOESS and GOES R Spacecraft this book prioritizes capabilities especially those related to climate research that were lost or placed at risk following the 2006 changes This book presents and recommends a prioritized short term strategy for recovery of crucial climate capabilities lost in the NPOESS and GOES R program descope However mitigation of these recent losses is only the first step in establishing a viable long term climate strategy one that builds on the lessons learned from the well intentioned but poorly executed merger of the nation s weather and climate observation systems **National Union Catalog** ,1982 Includes entries for maps and atlases **Commerce Business Daily** ,2000-06
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periodic table however because isotopes contain different numbers of neutrons each isotope has a unique atomic mass

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web aug 19 2022 atoms with the same number of protons but different numbers of neutrons are called isotopes they share almost the same chemical properties but differ in mass and therefore in physical properties there are stable isotopes which do not emit radiation and there are unstable isotopes which do emit radiation

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web may 16 2018 explanation an atom is characterized as a particular element by the number of positively charged massive nuclear particles i.e. Z . $Z=1$ we get hydrogen $Z=2$ we get helium $Z=3$ lithium $Z=47$ we get silver. The nucleus can also contain neutrons massive nuclear particles of zero charge.

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web atoms of the same element with different numbers of neutrons are called isotopes. Close isotopes are atoms of an element with the same number of protons and electrons but different numbers of neutrons.

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web oct 31 2023 an isotope is one of two or more species of atoms of a chemical element with the same atomic number and position in the periodic table and nearly identical chemical behavior but with different atomic masses and physical properties. Every chemical element has one or more isotopes.

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web jul 18 2018 answer link due to the number of neutrons isotopes differ with respect to the number of neutrons. Their nuclei contain where neutrons are massive particles of zero charge. A nucleus may contain variable numbers of neutrons and this gives rise to the phenomenon of isotopes.

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web isotopes are distinct nuclear species or nuclides. As a technical term of the same chemical element they have the same atomic number (number of protons in their nuclei) and position in the periodic table and hence belong to the same chemical element but differ in nucleon numbers (mass numbers) due to different numbers of neutrons.

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web verified by toppr correct option is d. Atoms having same atomic number but different atomic mass are known as isotopes. The difference in the mass is due to the difference in their number of neutrons with the same number of electrons and protons.

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web part i answer the questions based on the above reading. What is an isotope? Isotopes are versions of the same element. They have the same number of protons and electrons as the element but different mass numbers and number of neutrons. What does the number next to isotopes signify? The number indicates the isotope's mass number.

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web nov 27 2016 explanation basically isotopes are nuclei that have the same number of protons and are thus the same element but have unequal numbers of neutrons. Answer link the nucleus differs as a result of the number of neutrons in the atom. The actual structure doesn't differ all that much but its mass does differ. Thus its density is a little.

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