

Unit 1 Study Guide

Concept 1.1

1. Describe how genetic information is contained within DNA and is inherited from parent to offspring (See Figs 1.6-1.7) and explain is meant by "gene expression," (See Fig 1.8).

Concept 1.2

2. Explain the process of natural selection:
 - a. Describe the conditions that are necessary for natural selection to occur and explain why for each condition.
 - b. What causes mutations to occur and how does this relate to the conditions for natural selection?
 - c. Explain the consequences of natural selection (i.e., at what level does natural selection operate and at what level do we see evolutionary change? what is an adaptation?).
 - d. Apply your understanding of natural selection to a specific example to explain how and why this process can lead to different adaptations in different environments.

Concept 1.3

3. Differentiate between hypotheses and predictions and among different types of experimental variables (independent vs dependent vs controlled).
4. Interpret data presented in graphical form, drawing conclusions about the results as they relate to course themes (e.g., natural selection).

Concept 2.2

5. Describe the components of an atom (location, mass, and charge of protons, neutrons, and electrons) and how changes in each component affects atomic properties (e.g., what type of atom it is, whether it's an isotope, whether it is an ion, etc.)
6. What are valence electrons and how do they affect an atom's reactivity with other atoms?

Concept 2.3

7. Compare and contrast ionic vs covalent bonds.
8. Explain the difference between polar and nonpolar covalent bonds and how electronegativity affects covalent bond polarity.
9. When given examples of different bonds, predict bond polarity (nonpolar vs polar; rank bonds in order of polarity) and identify partial charges on a polar molecule.
10. Describe how hydrogen bonds form and contrast this with ionic and covalent bonds and Van der Waals interactions.
11. Explain how molecular shape can impact function.

Concept 2.5

12. Describe how the polarity of water molecules leads to the formation of hydrogen bonds between molecules of water (see Fig 2.16).

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Joseph Phillips



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