

Name: _____

Date: _____

Student Exploration: Nuclear Decay

Vocabulary: alpha particle, atomic number, beta particle, daughter product, gamma ray, isotope, mass number, nuclear decay, positron, radioactive, subatomic particle

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

The chart below gives the locations, charges, and approximate masses of three **subatomic particles**. The approximate mass of each particle is given in universal mass units (u).

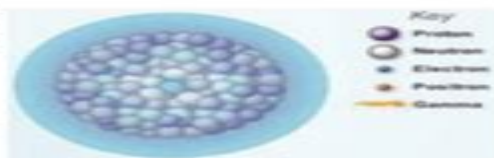
Particle	Location	Charge	Approximate mass
Proton	Nucleus	1^+	1 u
Neutron	Nucleus	0	1 u
Electron	Orbitals	1^-	0 u

- The **mass number** of an atom is equal to the sum of protons and neutrons in the nucleus. A helium atom has 2 protons and 2 neutrons. What is the mass number of this atom? _____
- The **atomic number** of an element is equal to the number of protons in each atom of the element. All helium atoms have 2 protons. What is the atomic number of helium? _____

Gizmo Warm-up

While most atoms are stable, some are **radioactive**, which means that they have a tendency to undergo spontaneous **nuclear decay**. The decay of radioactive atoms generally results in the emission of particles and/or energy.

Several types of nuclear decay can be explored with the Nuclear Decay Gizmo™. On the Gizmo, check that **Alpha decay** and **Uranium** are selected.



- Click **Play** (▶) and then click **Pause** (⏸) when the **alpha particle** is clearly visible. What is an alpha particle made of? _____
- Click **Play** and observe. Besides the alpha particle, what else is emitted from the nucleus during alpha decay? _____

Gamma rays are energetic electromagnetic waves; they are often emitted in nuclear decay.

Activity A: Alpha decay

Get the Gizmo ready:

- Click **Reset** (↺).
- Check that **Alpha decay** and **Uranium** are selected.



Question: How does alpha decay change the nucleus of a radioactive atom?

- Predict:** As you observed in the warm-up activity, an alpha particle consists of two protons and two neutrons. How will the emission of an alpha particle affect the following?
 - The atomic number of the atom: _____
 - The mass number of the atom: _____
- Calculate:** Turn on **Write equation**. What you see is an equation that shows the original uranium atom on the left. The boxes on the right represent the **daughter product**—the atom produced by radioactive decay—and the emitted alpha particle.
 - In the top left box, write the mass number of the daughter product and press "Enter" on your keyboard. What is this number? _____

Nuclear Decay Explore Learning Exploration Guide Answers

Michael Hoelscher



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1. Understanding the eBook Nuclear Decay Explore Learning Exploration Guide Answers
 - The Rise of Digital Reading Nuclear Decay Explore Learning Exploration Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Nuclear Decay Explore Learning Exploration Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nuclear Decay Explore Learning Exploration Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Nuclear Decay Explore Learning Exploration Guide Answers
 - Personalized Recommendations
 - Nuclear Decay Explore Learning Exploration Guide Answers User Reviews and Ratings
 - Nuclear Decay Explore Learning Exploration Guide Answers and Bestseller Lists

5. Accessing Nuclear Decay Explore Learning Exploration Guide Answers Free and Paid eBooks
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 - Nuclear Decay Explore Learning Exploration Guide Answers eBook Subscription Services
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6. Navigating Nuclear Decay Explore Learning Exploration Guide Answers eBook Formats
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 - Nuclear Decay Explore Learning Exploration Guide Answers Compatibility with Devices
 - Nuclear Decay Explore Learning Exploration Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Nuclear Decay Explore Learning Exploration Guide Answers
 - Highlighting and Note-Taking Nuclear Decay Explore Learning Exploration Guide Answers
 - Interactive Elements Nuclear Decay Explore Learning Exploration Guide Answers
8. Staying Engaged with Nuclear Decay Explore Learning Exploration Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Nuclear Decay Explore Learning Exploration Guide Answers
9. Balancing eBooks and Physical Books Nuclear Decay Explore Learning Exploration Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Nuclear Decay Explore Learning Exploration Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Nuclear Decay Explore Learning Exploration Guide Answers
 - Setting Reading Goals Nuclear Decay Explore Learning Exploration Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Nuclear Decay Explore Learning Exploration Guide Answers
 - Fact-Checking eBook Content of Nuclear Decay Explore Learning Exploration Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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