

Reactions of Amines and Epoxides

Preparation of Amines

Amination of alcohols: Polyalkylation is a serious problem for indirect synthesis of amines



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Bond Types, Strength, Length, Polarity and Dipole Moment

Two types of covalent bond

There are two types of covalent bond: bonding and non-bonding bond

- **Bonding bonds** are formed by the overlap of electron orbitals and are classified as **sigma bonds** (single bond) and **pi bonds** (double bond).



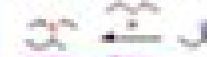
- **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Strength of Covalent Bonds



Bond and Bonding Energy



Bonding energy



There are two types of covalent bond

Definition of sigma bond



Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds

- **Sigma bonds** are formed by head-on overlap of orbitals.
- **Pi bonds** are formed by side-on overlap of orbitals.
- **Single bond** is a sigma bond.
- **Double bond** consists of one sigma and one pi bond.



pi bonds are weaker than sigma bonds

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar bond).
- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar bond).
- If the difference in electronegativity is greater than 1.7, the electrons are not shared (ionic bond).

- A difference of 1.7 or higher can be high for the electrons are not shared (ionic bond).

- A difference of 1.7 or higher can be high for the electrons are not shared (ionic bond).

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.
- The more electronegative atom has a partial negative charge (δ^-) and the less electronegative atom has a partial positive charge (δ^+).
- δ^- (delta minus) indicates an electron deficient atom.
- δ^+ (delta plus) indicates an electron rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moments

The polarity of the molecule is determined by the vector sum of the dipole moments. A polar molecule has a net dipole moment.



Consider the vector sum of the dipole moments for the molecule

Bond Length and Bond Strength

The bond length is the average distance between the nuclei of the two atoms.

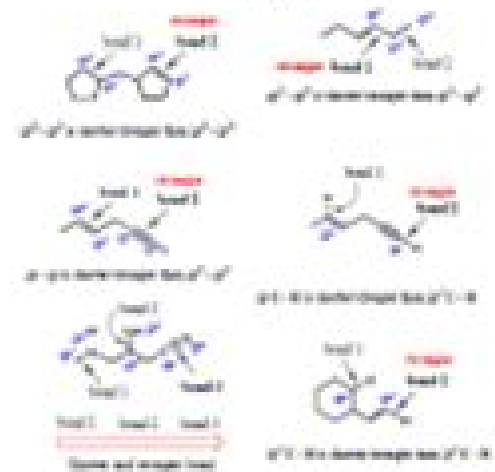


The bond length is the average distance between the nuclei of the two atoms.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3 (alkane)	sp^2-sp^2 (alkene)	$sp-sp$ (alkyne)
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	61 kcal/mol	83 kcal/mol
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
C-H Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.09 Å	1.08 Å	1.06 Å
Strength	99 kcal/mol	112 kcal/mol	120 kcal/mol

Higher s character = stronger and shorter bond

Example



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web terms in this set 11 a line on a map or chart along which there is a constant value isotherm a line on a map connecting points having the same temperature at a given time on average of a given period isobar a line on a map connecting points having the same atmospheric pressure at a given time on average of a given period

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web be interpreted as temperature therefore isotherm means same temperature isotherms drawn on temperature maps connect equal temperatures here are some helpful hints 1 isobars are drawn in 4 millibar intervals 2 always label the isobar value 3 pressures lower than the isobar value are always on one side of the isobar and higher