

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor. They are prepared by the reduction of nitriles.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Bond Types, Strength, Length, Polarity and Dipole Moment

Free-radical halogenation



There are two types of bonding: ionic bonding and covalent bonding.

- **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron).



- **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Electronegativity of Atoms



Bond and Electron Sharing



Electron sharing



There are two types of electronegativity

Electronegativity of atoms is



Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds.

- **Sigma bonds** are formed by head-on (end-to-end) overlap of orbitals.

- **pi bonds** are formed by side-on overlap of orbitals.

- **One single bond is a sigma bond.**

- **Two double bonds consist of one sigma and one pi bond.**



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity

values, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar bond). For example, the C-C and C-H bonds.

- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar bond). For example, the C-O and C=O bonds.

- A difference of 1.7 or higher is ionic (the electrons are not shared at all).

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is shown as a vector pointing from the less electronegative to the more electronegative atom.

- The vector electronegativity shows two important aspects: (1) the electronegativity of the atom, and (2) the electronegativity of the bond.

- **δ⁺** (delta plus) indicates an electron-deficient atom.

- **δ⁻** (delta minus) indicates an electron-rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moments

of the individual bonds is calculated. If the sum is not zero, the molecule is polar.



Overall the molecule is polar (net dipole moment is not zero).

Bond Length and Bond Strength

The longer the bond, the weaker it is.

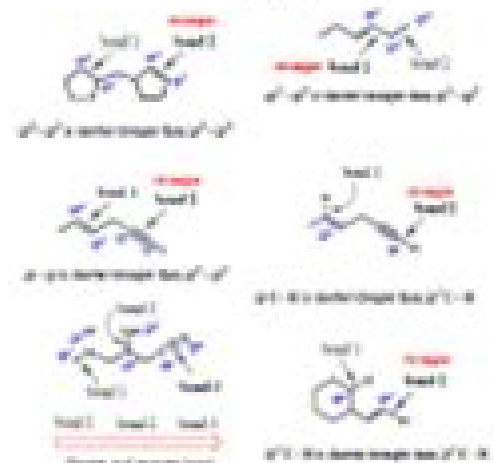


The longer the bond, the weaker it is.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.54 Å	1.34 Å	1.20 Å
Strength	376 kJ/mol	460 kJ/mol	505 kJ/mol
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
C-H Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.09 Å	1.08 Å	1.06 Å
Strength	413 kJ/mol	439 kJ/mol	505 kJ/mol

Higher s character = stronger and shorter bond

Examples



giving information on the strength of the bond.



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Arshad Iqbal



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