

# Reactions of Amines and Epoxides

## Preparation of Amines

Amines are often made by reduction of nitriles or by reductive amination of aldehydes and ketones.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

## Bond Types, Strength, Length, Polarity and Dipole Moment

### Two types of covalent bond

There are two types of covalent bond: **sigma bond** and **pi bond**.

- Sigma bonds** are formed by the overlap of atomic orbitals (AOs) in a head-on manner (end-to-end) and are **stronger**.



- Pi bonds** are formed by the overlap of atomic orbitals (AOs) in a side-on manner (side-to-side) and are **weaker**.



### The Relative Strength of $\sigma$ and $\pi$ Bonds



### Bond and Bond Energy



### Bond Energy



### There are two types of covalent bond

Definition of covalent bond



## Sigma and pi bonds

Covalent bonds are classified into **sigma** and **pi** bonds.

- Sigma bonds** are formed by the head-on overlap of atomic orbitals (AOs).
- Pi bonds** are formed by the side-on overlap of atomic orbitals (AOs).

- Sigma bonds** are **stronger** than pi bonds.
- Pi bonds** are **weaker** than sigma bonds.



$\sigma$  bonds are stronger than  $\pi$  bonds.

## Polar and Nonpolar Covalent Bonds

Depending on the difference in electronegativity, covalent bonds can be **polar** or **nonpolar**.

Electronegativity scale (Pauling scale):



- If the difference in electronegativity is **less than 0.5**, the bond is **nonpolar**.
- If the difference in electronegativity is **between 0.5 and 1.7**, the bond is **polar**.
- If the difference in electronegativity is **greater than 1.7**, the bond is **ionic**.

- A difference of **0.5 or higher** is **highly polar**.
- A difference of **1.7 or higher** is **highly ionic**.

## Dipole Moment

The polarity of a bond is described by the **dipole moment** ( $\mu$ ).



- The dipole moment is a vector pointing from the **less electronegative** atom to the **more electronegative** atom.
- The **size** of the dipole moment is a measure of the **charge separation**.
- $\mu$  (Debye)** is a unit of dipole moment.
- $\mu$  (Debye)** is a unit of dipole moment.

## Molecular Dipole

To determine the **molecular dipole**, the **vector sum** of the **bond dipoles** is used.

The **polarity** of the molecule is **indicated** by the **direction** of the **dipole moment**.



Sum the vector of the bond dipoles to find the molecular dipole.

## Bond Length and Bond Strength

The **length** of a bond is **related** to the **strength** of the bond.

| Bond     | H-H        | H-Cl       | H-Br       | H-I        |
|----------|------------|------------|------------|------------|
| Length   | 109 pm     | 127 pm     | 141 pm     | 154 pm     |
| Strength | 436 kJ/mol | 431 kJ/mol | 366 kJ/mol | 298 kJ/mol |

The **length** of a bond is **related** to the **strength** of the bond.

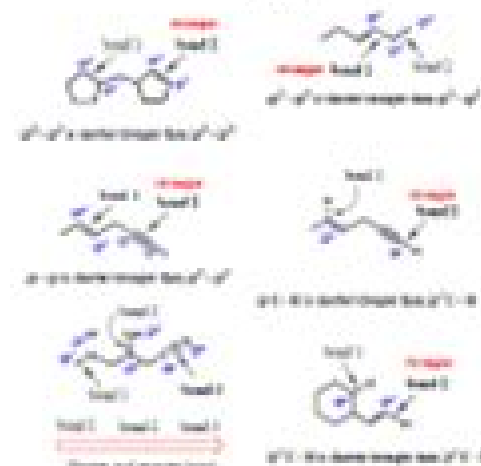


The **length** of a bond is **related** to the **strength** of the bond.

|               | H <sub>2</sub> -C≡C-H | H <sub>2</sub> -C≡C-H | H <sub>2</sub> -C≡C-H |
|---------------|-----------------------|-----------------------|-----------------------|
| Bond Length   | 109 pm                | 127 pm                | 141 pm                |
| Bond Strength | 436 kJ/mol            | 431 kJ/mol            | 366 kJ/mol            |
| Bond Angle    | 180°                  | 180°                  | 180°                  |
| Bond Energy   | 436 kJ/mol            | 431 kJ/mol            | 366 kJ/mol            |
| Bond Length   | 109 pm                | 127 pm                | 141 pm                |
| Bond Strength | 436 kJ/mol            | 431 kJ/mol            | 366 kJ/mol            |

Higher electronegativity difference leads to stronger bond.

## Examples



Stronger electronegativity difference leads to stronger bond.

# Organic Reactions Study Guide

**Kevin Smith**



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