

PHYSICS 10TH – Short / Detailed Question Answers

 NUCLEAR STRUCTURE

CHAPTER# 20

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Q.1: Why are some nuclei unstable?

Ans: Atomic nuclei consist of protons and neutrons, whereby protons repel each other through electrostatic force due to their positive charges. In contrast, nuclei bind the nucleons through another specific binding energy. These two forces compete with each other, leading to various nuclei stability. There are only certain neutron-proton pairings that form stable nuclei. As a result, an increasing ratio of neutrons to protons is required to form a stable nucleus. Some proportions of neutrons to protons are more stable than others in a nucleus if the neutron number, N , is plotted against the proton number, Z for all different isotopes of all the elements.

There are too many or few neutrons for a given number of protons. In that case, the resulting nucleus is unstable, and it undergoes radioactive decay to get rid of excess energy.

Q.2: Define the following terms:

- (i) Binding energy
- (ii) Higgs Bosons
- (iii) Radio-isotope OR Radioactive element
- (iv) Radioactivity
- (v) Gluon

Ans:

- (i) **Binding Energy:** Binding energy is the amount of energy required to separate a particle from a system of particles or to disperse all the particles of the system.
- (ii) **Higgs Bosons (Force Carrier):** It is a type of boson which carries mass as well and is known as the god particle.
- (iii) **Radio-isotope OR Radioactive Element:** If an isotope undergoes radioactive decay is called radio-isotope or radioactive element.
- (iv) **Radioactivity:** The emission of α , β and γ radiation with the release of energy is known as radioactivity.
- (v) **Gluon:** It is a type of boson (force carrier) that is an exchanging particle of strong nuclear force.

Nuclear Structure Study Guide Answers

D Keegan



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