

KEY

1. Complete and classify the following partial equations

- $2Fe + 6H_2SO_4 \rightarrow 2FeSO_4 + 3H_2$ (single rep)
- $2Fe + 3H_2SO_4 \rightarrow Fe_2(SO_4)_3 + 3H_2$
- $12s + 4Pb(NO_3)_2 \rightarrow 12s(PbO_2)_2 + 1Pb$ (single rep)
- $Al_2O_3 + HCl \rightarrow$
- $2Cu + O_2 \rightarrow 2CuO$ or $4Cu + 1O_2 \rightarrow 2Cu_2O$ (synthesis)
- $Al_2O_3 \rightarrow$
- $N_2 + 3H_2 \rightarrow 2NH_3$ (synthesis)
- $C_2H_6 + O_2 \rightarrow$
- $1Mg + 2HNO_3 \rightarrow 1Mg(NO_3)_2 + 1H_2$ (c.r.)
- $Na + H_2O \rightarrow$
- $3NaOH + H_3PO_4 \rightarrow Na_3PO_4 + 3H_2O$ (acid base)
- $H_2SO_4 + Al(OH)_3 \rightarrow$
- $2C_2H_6 + 7O_2 \rightarrow 4CO_2 + 6H_2O$ (combustion)
- $AlCl_3 \rightarrow$
- $2KOH + H_2SO_4 \rightarrow K_2SO_4 + 2H_2O$ (acid base)
- $C_2H_6 + O_2 \rightarrow$
- $1NaOH + HCl \rightarrow NaCl + 1H_2O$ (acid base)
- $NaHCO_3 + H_2SO_4 \rightarrow$
- $1C_2H_6 + 7O_2 \rightarrow 5CO_2 + 6H_2O$ (combustion)
- $H_2O \rightarrow$
- $2HNO_3 + Na_2CO_3 \rightarrow Na_2CO_3 + H_2O + 2HNO_3$ (c.r.)
- $CH_4 + O_2 \rightarrow$
- $1Pb + 4H_2SO_4 \rightarrow 1PbSO_4 + 3H_2$ (c.r.)
- $KI + Br_2 \rightarrow$
- $3Ba(OH)_2 + 2HNO_3 \rightarrow 1Ba(NO_3)_2 + 2H_2O$ (acid base)
- $Pb + Cu(NO_3)_2 \rightarrow$

2. Given: $2NaClO_3 \rightarrow 2NaCl + 3O_2$ calculate the mass of oxygen gas produced if 143 g of $NaClO_3$ is decomposed

65.4g

3. Calculate the mass of oxygen gas required to burn 242 g of wax ($C_{21}H_{44}$)

254g

4. If 14.0 moles of butane (C_4H_{10}) is burned

- calculate the number of moles of oxygen gas required (90.0 mol)
- calculate the number of moles of carbon dioxide produced (56.0 mol)
- calculate the mass of water produced (1080 g)

5. Metallic aluminum reacts with zinc nitrate to produce metallic zinc and aluminum nitrate

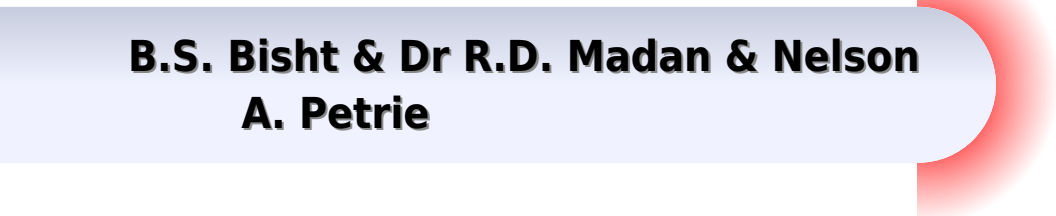
- Calculate the mass of aluminum that will react with 1.30 moles of zinc nitrate. (27.4g Al)
- Calculate the mass of zinc that would be produced from 50.0 g of aluminum and sufficient zinc nitrate. (182 g Zn)

6. Given: $Cu + 2AgNO_3 \rightarrow Cu(NO_3)_2 + 2Ag$

- How many grams of Cu are required to react with 3.00 g of silver nitrate? (0.544g Cu)
- How many grams of silver are produced from 3.00 g of silver nitrate and sufficient Cu? (1.70g Ag)

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**B.S. Bisht & Dr R.D. Madan & Nelson
A. Petrie**



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Burmeister, 2003 This invaluable book distils the research accomplishments of Professor Fred Basolo during the five decades when he served as a world leader in the modern renaissance of inorganic chemistry Its primary focus is on the very important area of chemistry known as coordination chemistry Most of the elements in the periodic table are metals and most of the chemistry of metals involves coordination chemistry This is the case in the currently significant areas of research including organometallic homogenous catalysis biological reactions of metalloproteins and even the solid state extended structures of new materials In these systems the metals are of primary importance because they are the sites of ligand substitution or redox reactions In the solid materials the coordination number of the metal and its stereochemistry are of major importance Some fifty years of research on transition metal complexes carried out in the laboratory of Professor Basolo at Northwestern University is recorded here as selected scientific publications The book is divided into three different

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