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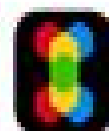
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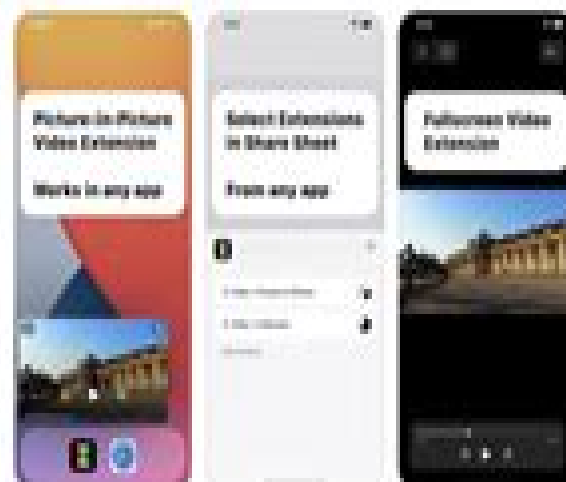


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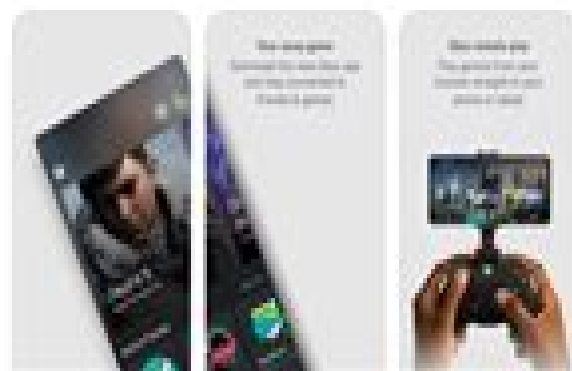


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web solution a titrant volume 0 mL the solution pH is due to the acid ionization of HCl because this is a strong acid the
ionization is complete and the hydronium ion molarity is 0.100 M the pH of the solution is then $pH = -\log 0.100 = 1.000$ pH $\log 0.$
100 1.000 b titrant volume 12.50 mL

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Volume of NaOH added (mL)	pH of the solution
4	6.4
6	6.5
9	5.9
5	5.9

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