Ways to Make the Most of Your Credit Card Offers



Credit Card Offers Guide

Clemens Wendtner

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web in today s experiment you will determine the solubility product ksp of calcium hydroxide ca oh 2 by measuring the concentration of ca oh 2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution ca oh 2 à ca 2 2oh 5 x 10 2 m calcium ions 5 x 10 2 m 2 0 m hydroxide ions calculate calcium hydroxide ksp using the concentrations of calcium ions and hydroxide ions

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web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the ksp of a slightly soluble basic salt and how solubility changes with temperature ksp is another form of equilibrium expression that deals with solubility the reaction that will take place is ca oh 2 ca2 2 oh

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web the chemical equation used was ca oh 2 s ca2 aq 2oh aq the solubility product was determined to be ksp ca2 oh from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

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web therefore it was determined through error calculations and ease of usethat the most reliable valid method to use to get the closest ksp value tothe theoretical ksp for ca was the determination through ph method 2 introduction calcium hydroxide is a strong base soluble in water

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web we will filter the solid calcium hydroxide from the solution and then titrate the dissolved calcium hydroxide in solution with a known concentration of hydrochloric acid using titrimetric calculations we will determine the ca 2 and oh concentrations in solution

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