

8. Identify the substances that are oxidised and the substances that are reduced in the following reactions.
 (a) $4\text{Na}(s) + \text{O}_2(g) \longrightarrow 2\text{Na}_2\text{O}(s)$ (b) $\text{CuO}(s) + 4\text{H}_2(g) \longrightarrow \text{Cu}(s) + \text{H}_2\text{O}(l)$

Ans. (a) In this reaction, sodium is oxidised to sodium oxide (Na_2O) and oxygen is reduced.
 (b) In this reaction, copper oxide (CuO) is reduced to copper and hydrogen (H_2) is oxidised to water.

B. CHAPTER-END EXERCISES

1. Which of the following statements about the reaction below are incorrect?



- (a) Lead is getting reduced. (b) Carbon dioxide is getting oxidised.
 (c) Carbon is getting oxidised. (d) Lead oxide is getting reduced.
 (i) (a) and (b) (ii) (a) and (c) (iii) (a), (b) and (c) (iv) all

Ans. The statements (a) and (b) are incorrect because the lead oxide is getting reduced and carbon is getting oxidised. Therefore, the answer is (i).

2. $\text{Fe}_2\text{O}_3 + 2\text{Al} \longrightarrow \text{Al}_2\text{O}_3 + 2\text{Fe}$

The above reaction is an example of a

- (a) Combination reaction
 (b) Double displacement reaction
 (c) Decomposition reaction
 (d) Displacement reaction

Ans. (d) The given reaction is a displacement reaction in which Fe of Fe_2O_3 is being replaced by Al.

3. What happens when dilute hydrochloric acid is added to iron filings? Choose the correct answer.

- (a) Hydrogen gas and iron chloride are produced.
 (b) Chlorine gas and iron hydroxide are produced.
 (c) No reaction takes place.
 (d) Iron, salt and water are produced.

Ans. (a) Hydrogen gas and iron chloride are produced according to the following reaction:



4. What is a balanced chemical equation? Why should chemical equations be balanced?

Ans. A balanced chemical equation is that which has same number of atoms of each element on both sides of the arrow.

The chemical equations are balanced to satisfy the law of conservation of mass according to which the total mass of all the elements present in the reactants must be equal to the total mass of all the elements present in the products. Therefore, the number of atoms of each element in the reactants must be equal to the number of atoms of these elements in the products.

5. Translate the following statements into chemical equations and then balance them.

- (a) Hydrogen gas combines with nitrogen to form ammonia.
 (b) Hydrogen sulphide gas burns in air to give water and sulphur dioxide.
 (c) Barium chloride reacts with aluminium sulphate to give aluminium chloride.
 (d) Potassium metal reacts with water to give potassium hydroxide and hydrogen gas.

Ans. (a) The skeleton equation for the reaction is given as:



Step 1. Count the number of atoms of each element on both sides of the equation.

Element	Number of atoms of each element of the reactants (LHS)	Number of atoms of each element of the products (RHS)
H	2	3
N	2	1

Step 2. In order to make the number of H atoms equal on both sides, put coefficient 3 before H_2 on LHS and coefficient 2 before NH_3 on the RHS.



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