

CHEMISTRY 1

FINAL EXAM REVIEW

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Name _____ Section _____ Date _____

Chemistry Review: Writing and Balancing Equations

Directions: Solve each problem and answer each question

1. Balance each equation below by adding the correct coefficients

- $___ \text{CuCl}_2 + ___ \text{NaBr} \rightarrow ___ \text{NaCl} + ___ \text{CuBr}_2$
- $___ \text{HCN} + ___ \text{Mg(OH)}_2 \rightarrow ___ \text{Mg(CN)}_2 + ___ \text{H}_2\text{O}$
- $___ \text{HNO}_3 + ___ (\text{NH}_4)_2\text{SO}_4 \rightarrow ___ \text{NH}_4\text{NO}_3 + ___ \text{H}_2\text{SO}_4$
- $___ \text{ZnCO}_3 + ___ \text{Al}_2\text{O}_3 \rightarrow ___ \text{ZnO} + ___ \text{Al}_2\text{CO}_3$
- $___ \text{C}_4\text{H}_{10} + ___ \text{O}_2 \rightarrow ___ \text{CO}_2 + ___ \text{H}_2\text{O}$

2. Write and balance each equation. Include phase labels

- Aqueous sodium hydroxide reacts with aqueous sulfuric acid to form aqueous sodium sulfate and water
- Solid iron (III) hydroxide and aqueous sodium carbonate are formed from the reaction of solid iron (III) carbonate and aqueous sodium hydroxide
- Solid aluminum reacts with aqueous hydrobromic acid to form aqueous aluminum bromide and hydrogen gas
- Gaseous ethane (dicarbon hexahydride) reacts with oxygen gas to form gaseous carbon dioxide and gaseous water
- Solid calcium phosphite reacts with aqueous copper (II) nitrate to form aqueous calcium nitrate and solid copper (II) phosphite

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