

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Free-radical halogenation



There are two types of bonding: ionic bonding and covalent bonding.

• **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron).



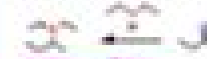
• **Covalent bonds** are formed by sharing of electrons between two atoms.



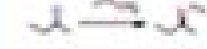
The Relative Electronegativity of Atoms



Bond and Bonding



Bonding energy



There are two types of bonding

Definition of ionic and covalent bonds



Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds.

• **Sigma bonds** are formed by head-on overlap of orbitals.

• **Pi bonds** are formed by side-on overlap of orbitals.

• **One single bond is a sigma bond.**

• **Two double bonds consist of one sigma and one pi bond.**



• **One bond is stronger than two bonds.**

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity of the atoms, covalent bonds can be polar and nonpolar.

• **Polar covalent bonds** are formed by the sharing of electrons between two atoms with different electronegativities.

• **Nonpolar covalent bonds** are formed by the sharing of electrons between two atoms with similar electronegativities.

• **If the difference in electronegativity is between 0.5 and 1.7, the bond is polar covalent.**

• **If the difference in electronegativity is between 1.7 and 2.0, the bond is ionic.**

• **A difference of 2.0 or higher indicates that the electrons are not shared but are lost to the more electronegative atom.**

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).

It is a vector quantity.



Electronegativity is a measure of the ability of an atom to attract electrons.

• The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.

• The more electronegative atom has a partial negative charge (δ^-) and the less electronegative atom has a partial positive charge (δ^+).

• **δ^- (delta minus) indicates an electron deficient atom.**

• **δ^+ (delta plus) indicates an electron rich atom.**

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar even if the dipoles are present. A good example is carbon dioxide.



Cancel the vector or the other pointing from the center out.

Bond Length and Bond Strength

The longer the bond, the weaker it is.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds and are shorter.

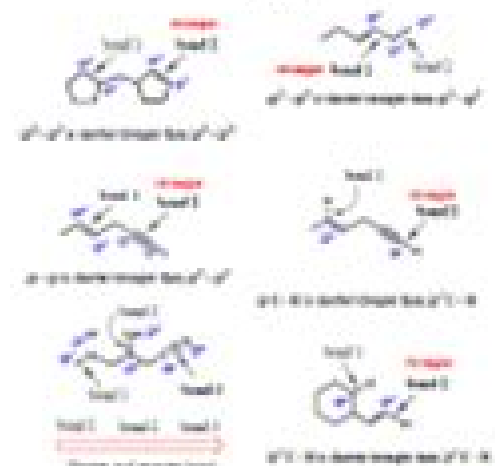


The longer the bond, the weaker it is.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3 - sp^3	sp^2 - sp^2	sp - sp
Hybridization	sp^3 - sp^3	sp^2 - sp^2	sp - sp
Length	154 pm	134 pm	121 pm
Strength	347 kJ/mol	461 kJ/mol	591 kJ/mol
Hybridization	sp^3 - sp^3	sp^2 - sp^2	sp - sp
C-H Bond	sp^3 - sp^3	sp^2 - sp^2	sp - sp
Length	109 pm	106 pm	103 pm
Strength	413 kJ/mol	460 kJ/mol	505 kJ/mol

Higher s character means shorter and stronger bond

Examples



giving information on the strength of the bond.



Organic Chemistry I Study Guide

Marc Loudon



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