

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor. They are prepared by the reduction of nitriles.



A Summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Epoxide and Nucleophile

Bond Types, Strength, Length, Polarity and Dipole Moment

Free-radical halogenation



There are two types of bonding: ionic bonding and covalent bonding.

• **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity. Examples:



• **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Electronegativity of Atoms



Bond and Electron Sharing



Bonding in Molecules



There are two types of electronegativity

Electronegativity of atoms is



Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds.

• **Sigma bonds** are formed by head-on overlap of orbitals.

• **pi bonds** are formed by side-on overlap of orbitals.

• **One single bond is a sigma bond.**

• **Two multiple bonds consist of one sigma and two pi bonds.**



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.

• **Polar bonds** are formed by the difference in electronegativity between two atoms.

• **Nonpolar bonds** are formed by the difference in electronegativity between two atoms.

• **If the difference in electronegativity is between 0.5 and 1.7, the bond is polar.**

• **If the difference in electronegativity is between 1.7 and 2.0, the bond is ionic.**

• **A difference of 2.0 or higher indicates that the electrons are not shared, but are lost to one atom.**

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).

It is a vector quantity.



It is a vector quantity.

• The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.

• The more electronegative atom has a partial negative charge (δ^-) and the less electronegative atom has a partial positive charge (δ^+).

• **δ^- (delta minus)** indicates an electron deficient atom.

• **δ^+ (delta plus)** indicates an electron rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moments is calculated.

The polarity of the molecule is determined by the vector sum of the dipole moments. A good example is carbon dioxide.



Carbon dioxide is a nonpolar molecule.

Bond Length and Bond Strength

The length of the bond is related to the strength of the bond.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds.

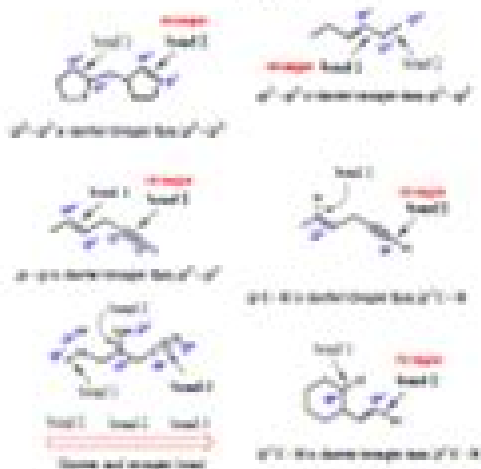


The length of the bond is related to the strength of the bond.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	154 pm	134 pm	121 pm
Strength	347 kJ/mol	461 kJ/mol	591 kJ/mol
Hybridization	sp^3-sp^2	sp^2-sp	$sp-sp$
Length	147 pm	133 pm	121 pm
Strength	355 kJ/mol	475 kJ/mol	591 kJ/mol

Higher s character means stronger bond

Examples



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