

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Types of Bonds

There are two types of bonds: ionic and covalent bonds.

- **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential and electron affinity).



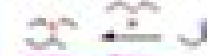
- **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Strength of Bonds



Bond and Bond Energy



Bond Energy



There are two types of bonds:

1. **Polar bonds**



There are two types of bonds:

2. **Nonpolar bonds**



There are two types of bonds:

3. **Coordinate covalent bonds**



Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is defined as a vector pointing from the less electronegative to the more electronegative atom.

- The vector dipole moment is a vector pointing from the less electronegative to the more electronegative atom.

- δ^- (delta minus) indicates an electron deficient atom.

- δ^+ (delta plus) indicates an electron rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar even if the dipoles are present. A good example is carbon dioxide.



Linear molecule with two polar bonds, but the dipoles cancel.

Bond Length and Bond Strength

The length of a bond is important.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds.

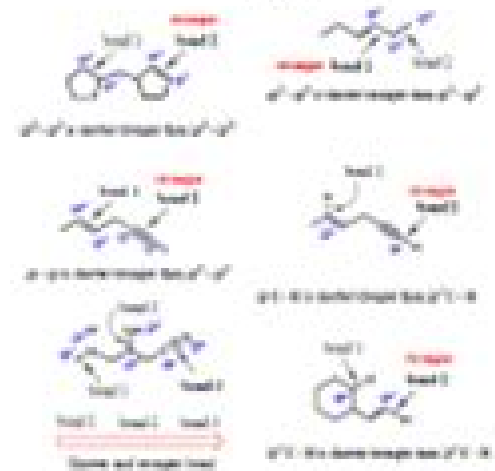


The length of a bond is important.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	154 pm	134 pm	121 pm
Strength	376 kJ/mol	460 kJ/mol	505 kJ/mol
Hybridization	sp^3-sp^2	sp^2-sp	$sp-sp$
Length	147 pm	133 pm	121 pm
Strength	410 kJ/mol	460 kJ/mol	505 kJ/mol

Higher s character means stronger bond

Examples



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Yasemin Kırkgöz, Kenan Dikilitaş



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