

Reactions of Amines and Epoxides

Preparation of Amines

Amines are often made by reduction of nitriles or by reductive amination of aldehydes and ketones.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Two types of covalent bond

There are two types of covalent bond: **sigma bond** and **pi bond**.

- Sigma bonds** are formed by the overlap of atomic orbitals (AOs) in a head-on manner (end-to-end overlap).



- Pi bonds** are formed by the overlap of atomic orbitals (AOs) in a side-on manner (side-to-side overlap).



The Relative Strength of σ and π Bonds



Bond and Bond Energy



Bond Energy



There are two types of covalent bond



Sigma and pi bonds

Covalent bonds are classified into **sigma** and **pi** bonds.

- Sigma bonds** are formed by the head-on overlap of atomic orbitals (AOs).
- Pi bonds** are formed by the side-on overlap of atomic orbitals (AOs).
- Sigma bonds** are stronger than pi bonds.
- Sigma bonds** are formed by the head-on overlap of atomic orbitals (AOs).



σ bonds are stronger than π bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in electronegativity, covalent bonds can be **polar** or **nonpolar**.



- If the difference in electronegativity is **less than 0.5**, the bond is **nonpolar**.
- If the difference in electronegativity is **between 0.5 and 1.7**, the bond is **polar**.
- If the difference in electronegativity is **greater than 1.7**, the bond is **ionic**.

- A difference of **0.5 or higher** is **highly polar**.
- A difference of **1.7 or higher** is **highly ionic**.

Dipole Moment

The polarity of a bond is described by the **dipole moment** (μ).



- The dipole moment is a vector pointing from the **less electronegative** atom to the **more electronegative** atom.
- The **net dipole moment** of a molecule is the vector sum of the individual bond dipole moments.
- Nonpolar** molecules have a net dipole moment of **zero**.
- Polar** molecules have a net dipole moment of **non-zero**.

Molecular Dipole

To determine the **molecular dipole**, the **vector sum** of the bond dipoles is used.

The **polarity** of the molecule is indicated by the **direction** of the net dipole moment.



Draw the vector sum of the bond dipoles to find the net dipole.

Bond Length and Bond Strength

The **length** of a bond is the **distance** between the two nuclei.

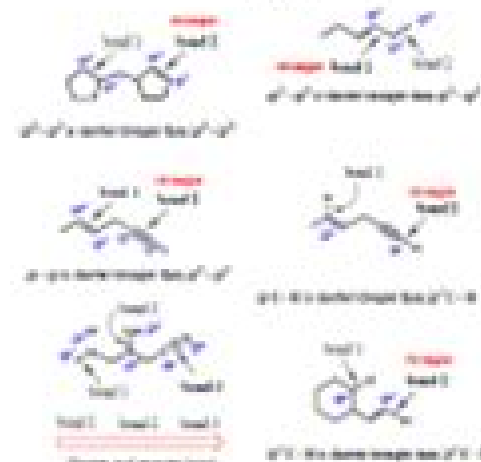


The **length** of a bond is the **distance** between the two nuclei.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	47 kcal/mol	58 kcal/mol
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
C-H Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.09 Å	1.06 Å	1.06 Å
Strength	99 kcal/mol	110 kcal/mol	120 kcal/mol

Higher s character = stronger and shorter bond

Examples



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web oct 9 2012 stephen king release date october 9 2012 source in the tall grass is a novella co written by stephen king and joe hill who had previously collaborated on the novella throttle first published in the june july 2012 issue of esquire the story was made available for download through the amazon kindle on 9 october 2012 summary

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