

Name: _____

ID#: _____

55 multiple-choice questions (3 pts each) + 6 fill-in questions (40 pts) = Total (205 pts).

1) How many **unpaired** electrons are present in the electron configuration of the **O** atom?

- A) none B) one
- C) two**
- D) three E) four

2) Which of the following contain(s) **polar covalent** bonds?

- A)
- H_2
- B)
- NaI
- C)
- Cl_2
- D) H_2O**
- E) both A and C

3) What is the formal charge and hybridization for the highlighted **Carbon**?

- A) 0,
- sp^2
- B) +1,
- sp^2
- C) -1, sp**
- D) +1,
- sp
- E) 0,
- sp

4) What of the following best describes the hybridization of C in **carbocation** CH_3^+ ?

- A)
- sp^3
- B) sp^2**
- C)
- sp
- D)
- s

5) what orbital is the nitrogen lone pair residing in the following molecule?

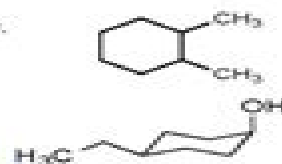
- A)
- sp^3
- B) sp^2**
- C)
- sp
- D)
- p
- E)
- s

6) Identify the **most acidic** carboxylic acid.**Answer is E**7) The **conjugate acid** of H_2O is _____

- A)
- HO^-
- B)
- H_3O
- C) H_3O^+**
- D)
- H_2O
- E)
- H_2O^+

8) How many potential stereoisomers exist for the following molecule.

- A) 1 B) 2
- C) 3**
- D) 4 E) 5



9) what is the name of the following compound?

- A)
- trans*
- 4-ethylcyclohexan-1-ol B)
- cis*
- 4-ethylhexan-1-ol
-
- C)
- trans*
- 4-methylcyclohexan-1-ol
- D) *cis*-4-ethylcyclohexan-1-ol**
-
- E)
- trans*
- 4-ethylcyclopentan-1-ol

10) which of the following step is part of the propagation in the methane chlorination?

- A)
- $Cl-Cl \longrightarrow 2 Cl \cdot$
- B)
- $H_3C \cdot + Cl-Cl \longrightarrow H_3C-Cl + Cl \cdot$
-
- C)
- $H_3C-H + \cdot Cl \longrightarrow H_3C \cdot + H-Cl$
- D) B and C

Answer is D

11) Based on the Hammond postulate, which of the following transition state best explains the following reaction step?

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Table of Contents Organic Chemistry 1 Final Exam Study Guide

1. Understanding the eBook Organic Chemistry 1 Final Exam Study Guide
 - The Rise of Digital Reading Organic Chemistry 1 Final Exam Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Organic Chemistry 1 Final Exam Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Organic Chemistry 1 Final Exam Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Organic Chemistry 1 Final Exam Study Guide
 - Personalized Recommendations
 - Organic Chemistry 1 Final Exam Study Guide User Reviews and Ratings
 - Organic Chemistry 1 Final Exam Study Guide and Bestseller Lists
5. Accessing Organic Chemistry 1 Final Exam Study Guide Free and Paid eBooks
 - Organic Chemistry 1 Final Exam Study Guide Public Domain eBooks
 - Organic Chemistry 1 Final Exam Study Guide eBook Subscription Services
 - Organic Chemistry 1 Final Exam Study Guide Budget-Friendly Options

6. Navigating Organic Chemistry 1 Final Exam Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Organic Chemistry 1 Final Exam Study Guide Compatibility with Devices
 - Organic Chemistry 1 Final Exam Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Organic Chemistry 1 Final Exam Study Guide
 - Highlighting and Note-Taking Organic Chemistry 1 Final Exam Study Guide
 - Interactive Elements Organic Chemistry 1 Final Exam Study Guide
8. Staying Engaged with Organic Chemistry 1 Final Exam Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Organic Chemistry 1 Final Exam Study Guide
9. Balancing eBooks and Physical Books Organic Chemistry 1 Final Exam Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Organic Chemistry 1 Final Exam Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Organic Chemistry 1 Final Exam Study Guide
 - Setting Reading Goals Organic Chemistry 1 Final Exam Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Organic Chemistry 1 Final Exam Study Guide
 - Fact-Checking eBook Content of Organic Chemistry 1 Final Exam Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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