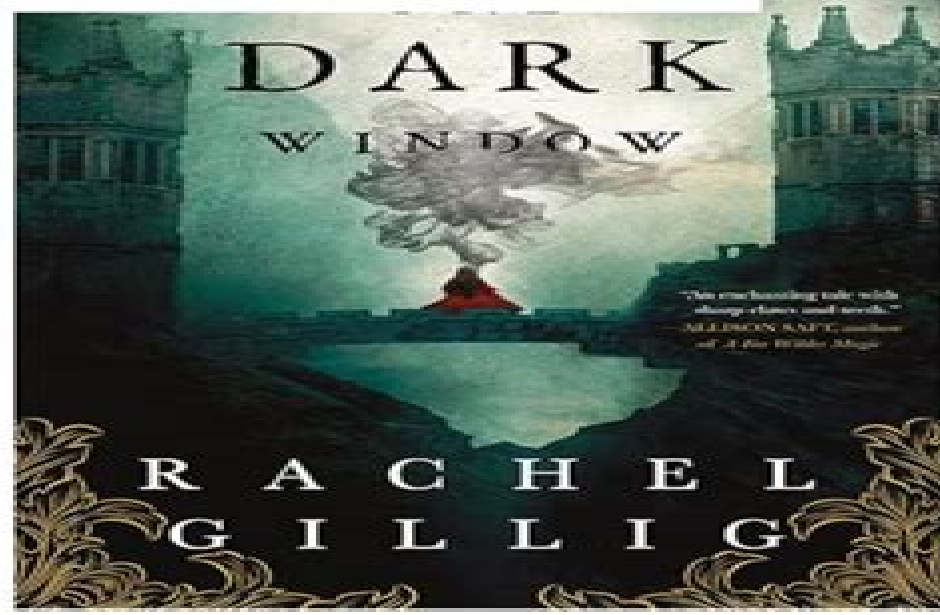
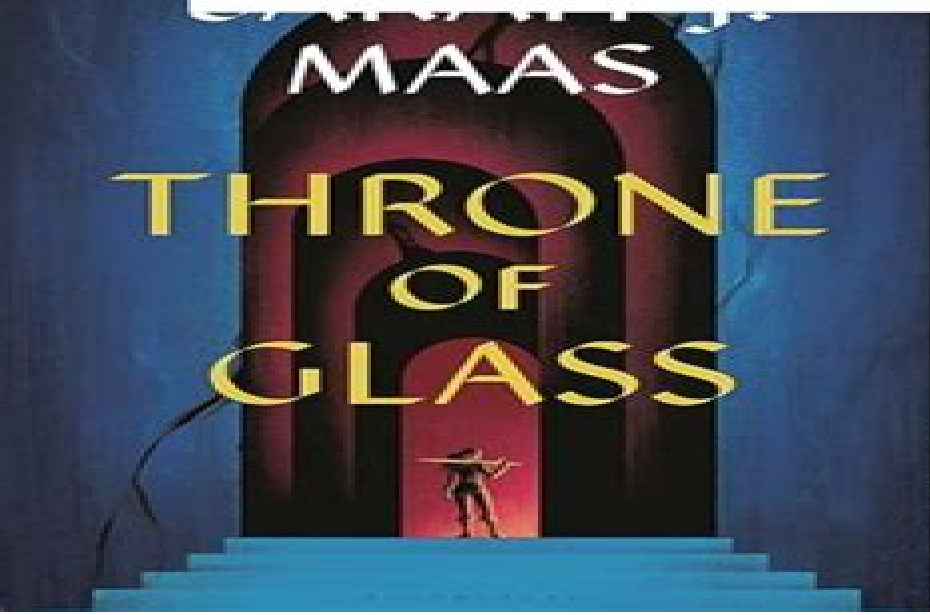




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web we utilize the equation that relate amount remaining initial mass and number of half lives $n = \frac{t}{T_{1/2}}$ $n = \frac{12}{24} = 0.5$ $125 \text{ g} \times 0.5 = 62.5 \text{ g}$ what is the half life of an isotope that is 80 remained after 16 days solution remaining 80 100 therefore $n = \frac{t}{T_{1/2}}$ $0.5 = \frac{16}{T_{1/2}}$ $T_{1/2} = \frac{16}{0.5} = 32$ now using the half life equation in example 1 we have

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