

Radiation, Nuclear Structure & Radioactivity - Review Answers

Conceptual Questions

1. What is the Duality Principle of electromagnetic radiation?

Answer: The duality principle states that electromagnetic radiation behaves both as a wave and a particle (photon). This explains phenomena like interference (wave) and the photoelectric effect (particle).

2. Why do photons with shorter wavelengths have higher energies?

Answer: Because energy is inversely proportional to wavelength ($E = hc/\lambda$). Shorter wavelengths mean higher frequency and thus higher energy.

3. What determines whether a nucleus is stable or unstable?

Answer: The balance between protons and neutrons in the nucleus. If this balance is off, the nucleus becomes unstable and undergoes radioactive decay.

4. How do electromagnetic waves transport energy without transporting matter?

Answer: They consist of oscillating electric and magnetic fields that carry energy through space without moving matter.

5. What are the key differences between alpha, beta, and gamma radiation in terms of penetration and biological risk?

Answer: Alpha: heavy, short range, high ionisation. Beta: lighter, moderate penetration. Gamma: deeply penetrating, can damage internal tissues.

6. Define the terms: isotope, isobar, isotone, and isomer. Give one example of each.

Answer: Isotope: same protons, diff neutrons (C-12, C-14). Isobar: same mass number (I-131, Xe-131). Isotone: same neutrons (e.g. 131I, 132Xe). Isomer: same nucleons, different energy states (99mTc, 99Tc).

7. What type of radiation is emitted by an excited nucleus returning to its ground state?

Answer: Gamma radiation (a high-energy photon).

Calculation Questions

1. Energy of 3.0 nm photon

Answer: Correct: 6.63×10^{-17} J

2. Convert 1.174 MeV to eV

Answer: Correct: 1.174×10^6 eV

3. Decay constant (10 hr half-life)

Answer: Correct: 0.0693 hr⁻¹

4. Half-life from decay (240 MBq to 15 MBq in 68 mins)

Nuclear Radiation Study Guide Answers

Morgan Lakeland



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