

Nuclear Chemistry - Guided Practice Questions**Multiple Choice** - Choose the one best answer and write its letter in the blank provided.

_____ 1) In balancing the nuclear reaction: $^{238}_{92}\text{U} \rightarrow ^{234}_{90}\text{Th} + \text{_____}$, the identity of the missing element is _____?

- A) Pu B) U C) Th D) Np E) Pa

_____ 2) The following rxn, $^{210}_{84}\text{Po} \rightarrow ^{210}_{82}\text{Pb} + \text{_____}$, is an example of what form of radioactive decay?

- A) alpha B) positron C) gamma D) beta E) e-capture

_____ 3) The missing product from this reaction is: $^{138}_{54}\text{Xe} \rightarrow ^{138}_{52}\text{Te} + \text{_____}$?

- A) $^0_{-1}\text{e}$ B) $^0_0\gamma$ C) ^4_2He D) ^1_0n E) ^1_1p

_____ 4) What is the mass number of a neutron _____?

- A) 1 B) 3 C) 2 D) 0 E) 4

_____ 5) In the nuclear decay reaction represented by $^{238}_{92}\text{U} \rightarrow ^{234}_{90}\text{Th}$, the emitted particle is _____.

- A) beta particle B) neutron C) alpha particle D) positron E) proton

_____ 6) The reaction shown below is responsible for creating C-14 in the atmosphere. What is the bombarding particle _____?



- A) electron B) proton C) positron D) alpha particle E) neutron

_____ 7) $^{131}_{53}\text{I}$ has a half-life of 8.04 days. Assuming you start with a 1.53 mg sample of $^{131}_{53}\text{I}$, how many mg will remain after 13.0 days _____?

- A) 0.440 B) 0.422 C) 0.268 D) 0.835 E) 0.499

_____ 8) All atoms of a given element have the same _____.

- A) atomic mass B) mass number C) # of nucleons D) atomic # E) # of neutrons

_____ 9) Atoms containing radioactive nuclei are called _____.

- A) radioisotopes B) radionuclides C) nuclides D) nucleons E) radionophores

_____ 10) What happens to the mass number and the atomic number of an element when it undergoes beta decay?

- A) The mass number decreases by 4 and the atomic number decreases by 2.
 B) The mass number does not change and the atomic number decreases by 2.
 C) Neither the mass number nor the atomic number change.
 D) The mass number increases by 2 and the atomic number increases by 1.
 E) The mass number does not change and the atomic number increases by 1.

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Scott C. Dulebohn



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Table of Contents Nuclear Chemistry Guided Answers

1. Understanding the eBook Nuclear Chemistry Guided Answers
 - The Rise of Digital Reading Nuclear Chemistry Guided Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Nuclear Chemistry Guided Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nuclear Chemistry Guided Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Nuclear Chemistry Guided Answers
 - Personalized Recommendations
 - Nuclear Chemistry Guided Answers User Reviews and Ratings
 - Nuclear Chemistry Guided Answers and Bestseller Lists
5. Accessing Nuclear Chemistry Guided Answers Free and Paid eBooks
 - Nuclear Chemistry Guided Answers Public Domain eBooks
 - Nuclear Chemistry Guided Answers eBook Subscription Services
 - Nuclear Chemistry Guided Answers Budget-Friendly Options

6. Navigating Nuclear Chemistry Guided Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Nuclear Chemistry Guided Answers Compatibility with Devices
 - Nuclear Chemistry Guided Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Nuclear Chemistry Guided Answers
 - Highlighting and Note-Taking Nuclear Chemistry Guided Answers
 - Interactive Elements Nuclear Chemistry Guided Answers
8. Staying Engaged with Nuclear Chemistry Guided Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Nuclear Chemistry Guided Answers
9. Balancing eBooks and Physical Books Nuclear Chemistry Guided Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Nuclear Chemistry Guided Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Nuclear Chemistry Guided Answers
 - Setting Reading Goals Nuclear Chemistry Guided Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Nuclear Chemistry Guided Answers
 - Fact-Checking eBook Content of Nuclear Chemistry Guided Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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