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 the concentrations of all species at equilibrium check your final answers with a q calculation let x be the number of mol l of h
 $2 \text{ or } i_2$ that is lost to reach equilibrium $h_2 \text{ g } i_2 \text{ g } 2 \text{ hi g}$

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web as a result of the reaction the equilibrium concentration of $\text{nh}_3 \text{ g}$ became 0 540 m what is the value of the equilibrium
 constant for this reaction at the given temperature of 700 c $n_2 \text{ g } 3h_2 \text{ g } 2\text{nh}_3 \text{ i}$ write the initial concentrations of reactants and
 products

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 represents the average kinetic energy of the molecules in that sample which gas variable p

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 at 425oc an equilibrium mixture has the following concentrations $hi_1 \text{ 01 x } 10^{-2} \text{ mol l } h_2 \text{ 1 25 x } 10^{-3} \text{ mol l } i_2 \text{ 1 49 x}$

10 3 mol l

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