

CUET Chemistry MCQs

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- (D) None of these
15. The number of tetrahedral and octahedral voids in hexagonal prismatic unit cell HCP is
(A) 8,4 (B) 2,1
(C) 12,6 (D) 6,12
16. Which of the following void is smallest in close packing?
(A) Tetrahedral (B) Octahedral
(C) Cubic (D) Square
17. The number of octahedral void in BCC structure is
(A) 0 (B) 1
(C) 2 (D) 4
18. Which of the following interstitial site is formed when the three closed packed spheres of one layer is put over three closed packed spheres of the second layer, their positions being inverted with respect to each other?
(A) Tetrahedral (B) Octahedral
(C) Rhombohedral (D) Tetragonal
19. In which crystal, alternate tetrahedral voids are occupied?
(A) NaCl (B) CaF₂
(C) ZnS (D) Na₂O
20. A metallic element crystallizes into a lattice containing a sequence of layers of ABABAB. Any packing of spheres leaves out voids in the lattice. Approximately what percentage by volume of this lattice is empty space?
(A) 74% (B) 26%
(C) 48% (D) 32%
21. Packing efficiency of an identical solid sphere is 74% in
(A) Simple cubic structure
(B) Monoclinic structure
(C) HCP structure
(D) BCC structure

22. Which of the following expression represents total volume of atoms present in a FCC unit cell of a metal? (r is the atomic radius)
(A) $\frac{20}{3}\pi r^3$ (B) $\frac{24}{3}\pi r^3$
(C) $\frac{12}{3}\pi r^3$ (D) $\frac{16}{3}\pi r^3$
23. In a BCC packing, the nearest neighbours lie along the
(A) edges of the cube
(B) face diagonal
(C) line joining the two opposite corners of the face.
(D) cube diagonal
24. The ionic radii of Rb⁺ and I⁻ are 1.46 and 2.16 Å, respectively. The most probable type of structure exhibited by it is
(A) CsCl type (B) NaCl type
(C) ZnS type (D) CaF₂ type
25. The radius of Ag⁺ ion is 126 pm while of I⁻ ion is 216 pm. The coordination number of Ag in AgI is
(A) 2 (B) 4
(C) 6 (D) 8
26. Structure shown here represents:

(A) Schottky defect
(B) Frenkel defect
(C) Metal excess defect
(D) None of these
27. Select the incorrect statement:
(A) Stoichiometry of crystal remains unaffected due to Schottky defect.

Multiple Choice Questions Of Solid State Chemistry

Peipei Pang



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Table of Contents Multiple Choice Questions Of Solid State Chemistry

1. Understanding the eBook Multiple Choice Questions Of Solid State Chemistry
 - The Rise of Digital Reading Multiple Choice Questions Of Solid State Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Questions Of Solid State Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Multiple Choice Questions Of Solid State Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Questions Of Solid State Chemistry
 - Personalized Recommendations
 - Multiple Choice Questions Of Solid State Chemistry User Reviews and Ratings
 - Multiple Choice Questions Of Solid State Chemistry and Bestseller Lists
5. Accessing Multiple Choice Questions Of Solid State Chemistry Free and Paid eBooks
 - Multiple Choice Questions Of Solid State Chemistry Public Domain eBooks
 - Multiple Choice Questions Of Solid State Chemistry eBook Subscription Services
 - Multiple Choice Questions Of Solid State Chemistry Budget-Friendly Options

6. Navigating Multiple Choice Questions Of Solid State Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Multiple Choice Questions Of Solid State Chemistry Compatibility with Devices
 - Multiple Choice Questions Of Solid State Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Multiple Choice Questions Of Solid State Chemistry
 - Highlighting and Note-Taking Multiple Choice Questions Of Solid State Chemistry
 - Interactive Elements Multiple Choice Questions Of Solid State Chemistry
8. Staying Engaged with Multiple Choice Questions Of Solid State Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Multiple Choice Questions Of Solid State Chemistry
9. Balancing eBooks and Physical Books Multiple Choice Questions Of Solid State Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Multiple Choice Questions Of Solid State Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Multiple Choice Questions Of Solid State Chemistry
 - Setting Reading Goals Multiple Choice Questions Of Solid State Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Multiple Choice Questions Of Solid State Chemistry
 - Fact-Checking eBook Content of Multiple Choice Questions Of Solid State Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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