

Reactions of Amines and Epoxides

Preparation of Amines

Amines are often made by reduction of nitriles or by Gabriel phthalimide synthesis



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Bond Types, Strength, Length, Polarity and Dipole Moment

Free-radical halogenation



There are two types of halogen bond: covalent and ionic.

- Covalent bonds are formed by the sharing of electrons from one orbital to another (homogeneous reaction) and ionic (heterogeneous) reactions.



- Covalent bonds are formed by sharing of electrons between two atoms.



The Relative Electronegativity of Atoms



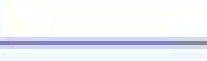
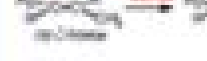
Bond and Reaction Rates



Bonding in Molecules



There are two types of bond: covalent and ionic.



Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds.

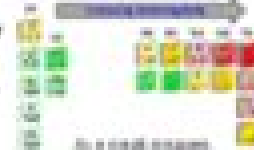
- Sigma bonds are formed by head-on overlap of orbitals.
- pi bonds are formed by side-on overlap of orbitals.
- A single bond is a sigma bond.
- A double bond consists of one sigma and one pi bond.



pi bonds are weaker than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar bond).
- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar bond).
- If the difference in electronegativity is greater than 1.7, the electrons are not shared (ionic bond).

A difference of 1.7 or higher can be high for the electrons are not shared (ionic bond) or a lower bond is formed.

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative to the more electronegative atom.
- The more electronegative atom has a partial negative charge (δ^-) and the less electronegative atom has a partial positive charge (δ^+).
- δ^- (delta minus) indicates an electron deficient atom.
- δ^+ (delta plus) indicates an electron rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar even if the dipoles are present. A good example is carbon dioxide.



Carbon dioxide is a linear molecule and the dipoles cancel out.

Bond Length and Bond Strength

The longer the bond, the weaker it is.

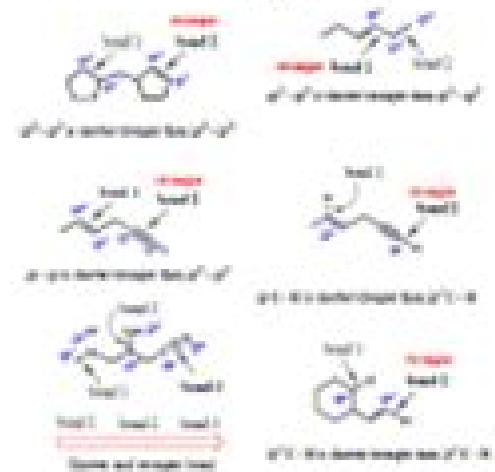


The longer the bond, the weaker it is.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	47 kcal/mol	58 kcal/mol
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
C-H Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.09 Å	1.06 Å	1.03 Å
Strength	98 kcal/mol	110 kcal/mol	120 kcal/mol

Higher s character = stronger and shorter bond

Examples



Organic Chemistry Mechanism Study Guide

Helga Völz, Christian Weiske



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of the authors approach to organic chemistry is to emphasize the relationship between structure and reactivity To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms The authors philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us

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Advanced Organic Chemistry - Structure and Mechanisms Mr. Ajeet Singh, 2018-05-22 Focused on advanced topics in organic chemistry this book explains structural theory stereochemistry and intricate reaction mechanisms equipping readers with the analytical

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Fryhle, Scott A. Snyder, 2023 Organic Chemistry 13th edition provides a comprehensive yet accessible treatment of all the essential organic chemistry concepts with emphasis on relationship between structure and reactivity in the subject The textbook includes all the concepts covered in a typical organic chemistry textbook but is unique in its skill development approach to the subject Numerous hands on activities and real world examples are integrated throughout the text to help students understand both the why and the how behind organic chemistry This International Adaptation offers new and updated content with improved presentation of all course material It offers new material on several topics including the relevance of intermolecular forces in the immune response and vaccines like those for Covid 19 the chemistry of breathing carbonic anhydrase how conjugation and complexation affect the color of lobsters and how biodegradable polymers are used to stabilize vaccines and pharmaceuticals Content is revised to reflect the current understanding of chemical processes and improved depictions of longstanding mechanisms This edition builds on the ongoing pedagogical strength of the book with the inclusion of additional worked and end of chapter problems and an engaging set of new problems entitled Chemical Consultant Needed These draw from the primary chemical literature and give students experience of working with more complex polyfunctional structures and areas where key transformations take place

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Chemistry of the Carbonyl Group Timothy K. Dickens, Stuart Warren, 2018-07-02 Teaches and enables students to build confidence in drawing and manipulating curly arrows a fundamental skill for all organic chemists This book is an interactive approach to learning about chemistry of the carbonyl group inviting students to work through its pages with pencil and paper in hand It educates with the belief that the most effective way to learn is by practice and interaction With this in mind the reader is asked to predict what would happen under a specific set of reaction conditions The book is divided into frames each frame poses a question and invites the reader to predict what will happen Subsequent frames give the solution but then pose more questions to develop a theme further Chemistry of the Carbonyl Group A Programmed Approach to Organic Reaction Mechanisms Revised Edition provides a solid grounding in the fundamental reactions of carbonyls Presented in full colour to enhance the understanding of mechanisms within chemistry the chapters of this step by step guide cover nucleophilic addition to the carbonyl group nucleophilic substitution nucleophilic substitution at the carbonyl group with complete removal of carbonyl oxygen carbanions and enolisation and building organic molecules from carbonyl compounds A must have book for undergraduate chemists to emphasise understanding in carbonyl group chemistry Goes through all the stages of basic carbonyl chemistry detailing even the simplest mechanisms A step by step learning guide to synthetic chemistry for the first year of a chemistry degree with all the information needed for independent learning Provides a solid grounding in the fundamental reactions of carbonyls which will inform the understanding of many other organic

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Table of Contents Organic Chemistry Mechanism Study Guide

1. Understanding the eBook Organic Chemistry Mechanism Study Guide
 - The Rise of Digital Reading Organic Chemistry Mechanism Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Organic Chemistry Mechanism Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Organic Chemistry Mechanism Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Organic Chemistry Mechanism Study Guide
 - Personalized Recommendations
 - Organic Chemistry Mechanism Study Guide User Reviews and Ratings
 - Organic Chemistry Mechanism Study Guide and Bestseller Lists
5. Accessing Organic Chemistry Mechanism Study Guide Free and Paid eBooks

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- 6. Navigating Organic Chemistry Mechanism Study Guide eBook Formats
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 - Organic Chemistry Mechanism Study Guide Compatibility with Devices
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- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Organic Chemistry Mechanism Study Guide
 - Highlighting and Note-Taking Organic Chemistry Mechanism Study Guide
 - Interactive Elements Organic Chemistry Mechanism Study Guide
- 8. Staying Engaged with Organic Chemistry Mechanism Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Organic Chemistry Mechanism Study Guide
- 9. Balancing eBooks and Physical Books Organic Chemistry Mechanism Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Organic Chemistry Mechanism Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Organic Chemistry Mechanism Study Guide
 - Setting Reading Goals Organic Chemistry Mechanism Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Organic Chemistry Mechanism Study Guide
 - Fact-Checking eBook Content of Organic Chemistry Mechanism Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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