ACS ORGANIC CHEMISTRY FINAL EXAM 2

- How can the free-radical addition reaction of
- HBr + alkene + peroxides
- 2. the polymerization of alkenes or dienesbe *initiated*?: heat, peroxides, or UV light
- What is the trend in stability for organic radicals/carbocations?: allylic/ben-zylic > 3 alkyl
 2 alkyl > 1 alkyl > methyl > vinylic/aryl

*Allylic and benzylic radicals are especially easy to generate because of the stabi-lization caused by resonance.

* when formed in a ring, the C+ will form preferentially to the C that can be stabilized by resonance.

- What conditions will usually indicate a radical reaction?: AIBN (a radicalinitiator), peroxides, UV light, or high temps (300-500 C)
- Groups that reduce electron density of a ring by resonance or inductionwill: cause the ring to be less reactive than benzene
- Groups that increase the electron density of the ring by resonance orinduction will: be more reactive than benzene
- Electron-donating groups cause the electron density to increase at the: or-tho (next to) and para (opposite) positions. Causes electrophiles to prefer to attackat those positions
- Electron withdrawing groups cause electron density to be reduced at the: ortho and para
 position more than at the meta position. Causes electrophilesto prefer to attack at the meta
 position.
- What has the greatest influence on electron density of an aromatic ring?-

: Resonance

- What increases electron density in an aromatic ring?: Groups that have anatom with nonbonded electrons
- What reduces electron density?: groups that have multiple bonds to an atommore electronegative than carbon
- What are the two distinct pathways for nucleophilic aromatic substitution?-

: Addition-elimination and elimination-addition.

· What is the addition-elimination mechanism?: 1. aromatic ring accepts a pair of electrons

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American Society for Engineering Education

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