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ACT English



68 Concepts

95 Quizzes

3

Correcting Grammatical Errors

00:15:24

3

Phrase, Clause, and Sentence Errors

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00:08:57

4

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00:00:36

Subordinating Conjunction Errors

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water $n_a = 54 \text{ g} / 18 \text{ g mol}^{-1} = 3 \text{ mol}$ total number of moles $n_a + n_b = 0.5 + 3 = 3.5 \text{ mol}$ mole fraction of solute ethyl alcohol $x_b = n_b / n_a + n_b = 0.5 / 3.5 = 0.1429$

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40 g of H_2O molar mass of H_2O is 18 gm mole and molar mass of CH_3OH is 32 gm mole solution moles of CH_3OH $4.5 / 32 = 0.14$

mole moles of H_2O 40.1822 moles therefore according to the equation

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time the total pressure $p_a \times \text{mole fraction}_a = \text{mole}_a / \text{moles total}$ right sum $n_1 \text{ all } x_{n_1}$ mole fraction of substance
a is often denoted by x_a and

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present is equal to one $x_a + x_b + \dots = 1$ when the mole fractions are multiplied by 100 they give the mole percentage mole fraction is
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12.01 + 4(35.45) = 153.84 g/mol

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g of water calculate the molarity and molality of the solution given $n = 14$ $h = 1$ $c = 12$ $o = 16$ given mass of solute urea 11.11 g mass
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