

Section 7.6: Solubility Equilibria and the Solubility Product Constant

Tutorial 1 Practice, page 462

1. **Solution:** $[\text{MgCO}_3(\text{s})]$ is constant, so it is not included in the equation. Therefore,
 $K_{\text{sp}} = [\text{Mg}^{2+}(\text{aq})][\text{CO}_3^{2-}(\text{aq})]$

Statement: The solubility product constant equation for the solubility equilibrium is
 $K_{\text{sp}} = [\text{Mg}^{2+}(\text{aq})][\text{CO}_3^{2-}(\text{aq})]$

2. **Given:** iron(III) hydroxide dissolves in water

Required: K_{sp} equation

Solution:

Step 1. Write a balanced equation for the dissolution of iron(III) hydroxide in water.



Step 2. $[\text{Fe}(\text{OH})_3(\text{s})]$ is constant, so it is not included in the equation. Therefore,

$$K_{\text{sp}} = [\text{Fe}^3+(\text{aq})][\text{OH}^-(\text{aq})]^3$$

Statement: The solubility product constant equation for iron(III) hydroxide is

$$K_{\text{sp}} = [\text{Fe}^3+(\text{aq})][\text{OH}^-(\text{aq})]^3$$

3. **Given:** At 25 °C: $[\text{Ca}^{2+}(\text{aq})] = 4.53 \times 10^{-7} \text{ mol/L}$; $[\text{PO}_4^{3-}(\text{aq})] = 3.02 \times 10^{-7} \text{ mol/L}$

Required: K_{sp} of calcium phosphate at 25 °C

Solution:

Step 1. Write a balanced equation for the solution equilibrium.



Step 2. Write the solubility product constant equation for the equilibrium, omitting any solids.

$$K_{\text{sp}} = [\text{Ca}^{2+}(\text{aq})]^3[\text{PO}_4^{3-}(\text{aq})]^2$$

Step 3. Substitute the given ion concentrations into the K_{sp} equation and solve.

$$\begin{aligned} K_{\text{sp}} &= [\text{Ca}^{2+}(\text{aq})]^3[\text{PO}_4^{3-}(\text{aq})]^2 \\ &= (4.53 \times 10^{-7})^3(3.02 \times 10^{-7})^2 \\ K_{\text{sp}} &= 8.48 \times 10^{-33} \end{aligned}$$

Statement: The K_{sp} of calcium phosphate in a saturated solution at 25 °C is 8.48×10^{-33} .

Tutorial 2 Practice, page 464

1. **Given:** solubility of $\text{MgF}_2(\text{s}) = 1.72 \times 10^{-3} \text{ g/100 mL}$ at 25 °C

Required: K_{sp} of $\text{MgF}_2(\text{s})$ at 25 °C

Solution:

Step 1. Write the balanced equation.



Step 2. Write the solubility product constant equation.

$$K_{\text{sp}} = [\text{Mg}^{2+}(\text{aq})][\text{F}^-(\text{aq})]^2$$

Step 3. Determine the equilibrium concentrations in mol/L of the dissolved ions. Use an ICE table. Since the concentration of solid MgF_2 remains constant, it is not included.

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 - The Rise of Digital Reading Nelson 12 Solution Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Nelson 12 Solution Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nelson 12 Solution Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Nelson 12 Solution Manual
 - Personalized Recommendations
 - Nelson 12 Solution Manual User Reviews and Ratings
 - Nelson 12 Solution Manual and Bestseller Lists
5. Accessing Nelson 12 Solution Manual Free and Paid eBooks

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 - Highlighting and Note-Taking Nelson 12 Solution Manual
 - Interactive Elements Nelson 12 Solution Manual
- 8. Staying Engaged with Nelson 12 Solution Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Nelson 12 Solution Manual
- 9. Balancing eBooks and Physical Books Nelson 12 Solution Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Nelson 12 Solution Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Nelson 12 Solution Manual
 - Setting Reading Goals Nelson 12 Solution Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Nelson 12 Solution Manual
 - Fact-Checking eBook Content of Nelson 12 Solution Manual
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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