

CHEMISTRY 1

FINAL EXAM REVIEW

Table of Contents

30	Legal Stuff
40	Writing and Balancing Equations
50	Writing Compound Names and Formulas
60	Moles and Avogadro's Number
70	Stoichiometry
80	Empirical and Molecular Formula
90	The Periodic Table
100	Gas Law Calculations
110	Molarity and Dilution
120	Dimensional Analysis and Unit Conversions
130	Drawing Lewis Structures
140	Phase Diagrams and Phase Transitions
150	Electron Configurations
160	Intermolecular Forces
170	Periodic Trends
180	Significant Figures
190	Scientific Notation
200	Thermochemistry: Calorimetry and Enthalpy
210	Acids, Bases, and pH
220	Electrons and Quantum Numbers
230	Classifying matter, phase changes, and properties
240	Types of Chemical Reactions, Activity Series, and Solubility
25-440	Answers

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Name _____ Section _____ Date _____

Chemistry Review: Writing and Balancing Equations

Directions: Solve each problem and answer each question

1. Balance each equation below by adding the correct coefficients

- $___ \text{CuCl}_2 + ___ \text{NaBr} \rightarrow ___ \text{NaCl} + ___ \text{CuBr}_2$
- $___ \text{HCN} + ___ \text{Mg(OH)}_2 \rightarrow ___ \text{Mg(CN)}_2 + ___ \text{H}_2\text{O}$
- $___ \text{HNO}_3 + ___ (\text{NH}_4)_2\text{SO}_4 \rightarrow ___ \text{NH}_4\text{NO}_3 + ___ \text{H}_2\text{SO}_4$
- $___ \text{ZnCO}_3 + ___ \text{Al}_2\text{O}_3 \rightarrow ___ \text{ZnO} + ___ \text{Al}_2\text{CO}_3$
- $___ \text{C}_4\text{H}_{10} + ___ \text{O}_2 \rightarrow ___ \text{CO}_2 + ___ \text{H}_2\text{O}$

2. Write and balance each equation. Include phase labels

- Aqueous sodium hydroxide reacts with aqueous sulfuric acid to form aqueous sodium sulfate and water
- Solid iron (III) hydroxide and aqueous sodium carbonate are formed from the reaction of solid iron (III) carbonate and aqueous sodium hydroxide
- Solid aluminum reacts with aqueous hydrobromic acid to form aqueous aluminum bromide and hydrogen gas
- Gaseous ethane (dicarbon hexahydride) reacts with oxygen gas to form gaseous carbon dioxide and gaseous water
- Solid calcium phosphite reacts with aqueous copper (II) nitrate to form aqueous calcium nitrate and solid copper (II) phosphite

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1. Understanding the eBook Organic Chemistry 1 Klein Final Exam
 - The Rise of Digital Reading Organic Chemistry 1 Klein Final Exam
 - Advantages of eBooks Over Traditional Books
2. Identifying Organic Chemistry 1 Klein Final Exam
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Organic Chemistry 1 Klein Final Exam
 - User-Friendly Interface
4. Exploring eBook Recommendations from Organic Chemistry 1 Klein Final Exam
 - Personalized Recommendations
 - Organic Chemistry 1 Klein Final Exam User Reviews and Ratings
 - Organic Chemistry 1 Klein Final Exam and Bestseller Lists

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 - Highlighting and Note-Taking Organic Chemistry 1 Klein Final Exam
 - Interactive Elements Organic Chemistry 1 Klein Final Exam
8. Staying Engaged with Organic Chemistry 1 Klein Final Exam
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Organic Chemistry 1 Klein Final Exam
9. Balancing eBooks and Physical Books Organic Chemistry 1 Klein Final Exam
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Organic Chemistry 1 Klein Final Exam
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Organic Chemistry 1 Klein Final Exam
 - Setting Reading Goals Organic Chemistry 1 Klein Final Exam
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Organic Chemistry 1 Klein Final Exam
 - Fact-Checking eBook Content of Organic Chemistry 1 Klein Final Exam
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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