

Organic Chemistry, Second Edition

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Chapter 2 Lecture Outline

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2.1 Bronsted-Lowry Acids and Bases

- A Bronsted-Lowry acid is a proton donor.
- A Bronsted-Lowry base is a proton acceptor.
- H^+ = proton

Figure 2.1
Examples of Bronsted-Lowry
acids and bases.

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Bronsted-Lowry acids [H-A]		Bronsted-Lowry bases [B:]	
Inorganic	Organic	Inorganic	Organic
HO H_2SO_4 HSO_4^- H_2O H_3O^+	CH_3CO_2H acetic acid $HO_2CCH_2-C(OH)(COOH)-CH_2CO_2H$ citric acid	$H_2O:$ OH^- PH_3 SiH_4	CH_3NH_2 methylamine CH_3O^- methoxide $CH_3-C(=O)-CH_3$ acetone $CH_2=CH_2$ ethylene
- All Bronsted-Lowry acids contain a proton. - The net charge may be zero, (+), or (-).		- All Bronsted-Lowry bases contain a lone pair of electrons or a π bond. - The net charge may be zero or (-).	

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Organic Chemistry Second Edition Answers To Problems

Eric Tyson



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Chambers, 2018-11-13 The perfect companion for iMac users You've got a great iMac but now it's time to take things to the next level This easy to use guide shows you how to do stuff like setting up and configuring your iMac along with fun stuff like using the Photos Messages and iMovie apps You'll also discover how to get down to business with productivity apps like Numbers Pages and Keynote Finally you'll find help on how to sync everything to iCloud and stay organized with Family Sharing Personalize your iMac and find your way around macOS X Surf the web send Reminders and make FaceTime video calls Store your photos and files with iCloud Discover tricks to become more efficient iMac newbies won't want to be without this easy to follow companion guide

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