

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Types of Bonds

There are two types of bonds: ionic and covalent bonds.

- Ionic bonds are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron gain).



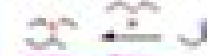
- Covalent bonds are formed by sharing of electrons between two atoms.



The Relative Strength of Bonds



Bond and Bond Energy



Bond Energy



There are two types of bonds

Definition of ionic bond



Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds.

- Sigma bonds are formed by head-on overlap of orbitals.
- pi bonds are formed by side-on overlap of orbitals.
- A single bond is a sigma bond.
- A double bond consists of one sigma and one pi bond.



pi bonds are weaker than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar bond).
- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar bond).
- If the difference in electronegativity is greater than 1.7, the electrons are not shared (ionic bond).

A difference of 1.7 or higher can be used to distinguish between ionic and covalent bonds.

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Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.
- The vector dipole moment is a product of the charge difference and the distance between the charges.
- $\mu = q \times r$ (where q is the charge difference and r is the distance between the charges).
- μ (dipole moment) indicates the direction of the bond.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moments is used.

The polarity of the molecule is determined by the vector sum of the dipole moments of the bonds. A good example is carbon dioxide.



Carbon dioxide is a nonpolar molecule.

Bond Length and Bond Strength

The bond length is the distance between the nuclei of the two atoms.



Double bonds are stronger than single bonds.

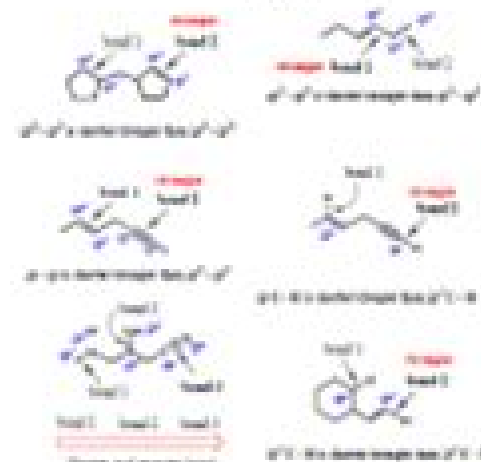


The bond length is the distance between the nuclei of the two atoms.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3 - sp^3	sp^2 - sp^2	sp - sp
Hybridization	sp^3 - sp^3	sp^2 - sp^2	sp - sp
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	47 kcal/mol	58 kcal/mol
Hybridization	sp^3 - sp^3	sp^2 - sp^2	sp - sp
C-H Bond	sp^3 - sp^3	sp^2 - sp^2	sp - sp
Length	1.09 Å	1.06 Å	1.03 Å
Strength	99 kcal/mol	112 kcal/mol	120 kcal/mol

Higher s character means stronger bond

Examples



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