

Reactions of Amines and Epoxides

Preparation of Amines

Amination of alcohols: Polyalkylation is a serious problem for indirect synthesis of amines



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Two different types of bonding

There are two types of bonding: ionic bonding and covalent bonding

- Ionic bonds are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron).



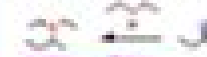
- Covalent bonds are formed by sharing of electrons between two atoms.



The Relative Electronegativity of an



Bonding and Bonding in a bond



Bonding in a bond



There are two types of bonding



Sigma and pi bonds

Covalent bonds are classified in sigma and pi bonds

- Sigma bonds are formed by head-on overlap of orbitals.
- pi bonds are formed by side-on overlap of orbitals.
- A single bond is a sigma bond.
- A double bond consists of one sigma and one pi bond.



pi bonds are stronger than sigma bonds

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar bond).
- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar bond).
- If the difference in electronegativity is greater than 1.7, the electrons are not shared (ionic bond).

- A difference of 1.7 or higher can be high for the electrons are not shared (ionic bond).

- A difference of 1.7 or higher can be high for the electrons are not shared (ionic bond).

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative to the more electronegative atom.
- The more electronegative atom has a partial negative charge (δ^-) and the less electronegative atom has a partial positive charge (δ^+).
- δ^- (delta minus) indicates an electron deficient atom.
- δ^+ (delta plus) indicates an electron rich atom.

Molecular Dipole

To increase the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar (even for dipole moments are cancelled). A good example is carbon dioxide.



Cancel the vector or the other pointing point in the same way

Bond Length and Bond Strength

The longer the bond, the weaker the bond



Double bonds are stronger than single bonds in the same row and a stronger bond

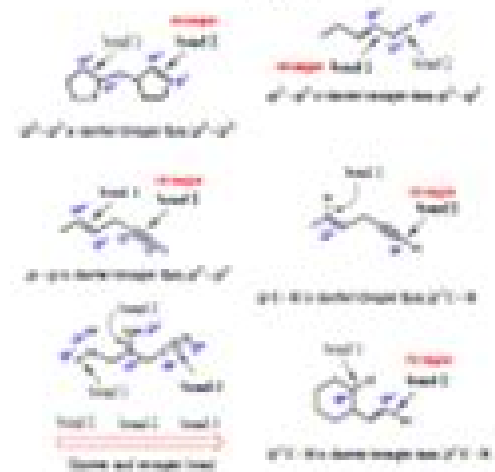


The longer the bond, the weaker the bond

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	154 pm	134 pm	121 pm
Strength	347 kJ/mol	461 kJ/mol	591 kJ/mol
Hybridization	sp^3-sp^2	sp^2-sp	$sp-sp$
C-H Bond	109 pm	108 pm	106 pm
Length	154 pm	134 pm	121 pm
Strength	347 kJ/mol	461 kJ/mol	591 kJ/mol

Higher s character = stronger and shorter bond

Examples



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J Rink



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