

Reactions of Amines and Epoxides

Preparation of Amines

Amines are often made by reduction of nitriles or by reductive amination of aldehydes and ketones.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Two types of covalent bond

There are two types of covalent bond: **sigma bond** and **pi bond**.

- Sigma bonds** are formed by the overlap of atomic orbitals (AOs) in a head-on manner (end-to-end).
- Pi bonds** are formed by the overlap of atomic orbitals (AOs) in a side-on manner (side-to-side).



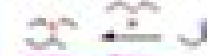
- Carbon bonds** are formed by sharing of electrons between two atoms.



The Relative Strength of σ and π Bonds



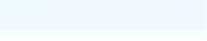
Bond and Bonding in a Molecule



Bonding in a Molecule



There are two types of covalent bond



Sigma and pi bonds

Carbon bonds are classified as **sigma** and **pi** bonds.

- Sigma bonds** are formed by head-on overlap of orbitals.
- Pi bonds** are formed by side-on overlap of orbitals.
- Sigma bonds** are stronger than pi bonds.
- Sigma bonds** are formed by the overlap of orbitals.



σ bonds are stronger than π bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in electronegativity, covalent bonds can be **polar** and **nonpolar**.



- If the difference in electronegativity is **less than 0.5**, the bond is **nonpolar**.
- If the difference in electronegativity is **between 0.5 and 1.7**, the bond is **polar**.
- If the difference in electronegativity is **greater than 1.7**, the bond is **ionic**.

- A difference of **0.5 or higher** is enough for the electron to be pulled toward the more electronegative atom.

- A difference of **1.7 or higher** is enough for the electron to be pulled toward the more electronegative atom.

Dipole Moment

The polarity of a bond is described by the **dipole moment** (μ).



- The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.
- The more electronegative atom has a partial negative charge (δ^-), and the less electronegative atom has a partial positive charge (δ^+).
- δ^- (delta minus) indicates an electron deficient atom.
- δ^+ (delta plus) indicates an electron rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The polarity of the molecule is determined by the vector sum of the dipoles. If the sum is zero, the molecule is nonpolar. If the sum is not zero, the molecule is polar.



Cancel the vector or the other pointing point in the same way.

Bond Length and Bond Strength

The length of the bond is important.



Double bonds are stronger than single bonds.

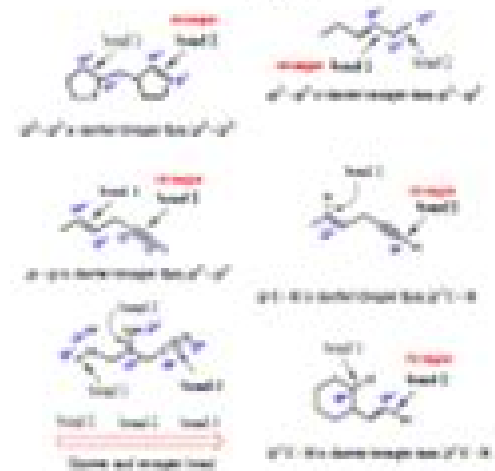


The length of the bond is important.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.54 Å	1.34 Å	1.20 Å
Strength	376 kJ/mol	472 kJ/mol	591 kJ/mol
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
C-H Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.54 Å	1.34 Å	1.20 Å
Strength	413 kJ/mol	472 kJ/mol	591 kJ/mol

Higher s character means stronger bond

Examples



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**T. W. Graham Solomons, Craig B.
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