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web oxidation reduction redox reactions oxidation $\text{Cu} \rightarrow \text{Cu}^{2+} + 2\text{e}^-$ reduction $\text{MnO}_4^- + 4\text{H}^+ + 3\text{e}^- \rightarrow \text{MnO}_2 + 2\text{H}_2\text{O}$ l when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients $\text{Cu} : \text{MnO}_4^-$ is 4 : 1

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^- + \text{Au} \rightarrow \text{AuCN}_2^- + 2\text{e}^-$ $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ make electrons equal convert to basic solution $4\text{CN}^- + 2\text{Au} \rightarrow 2\text{AuCN}_2^- + 2\text{e}^-$ multiplied by a factor of 2 $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ 3 the final answer

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the loss of electrons resulting in a more positively charged species reduction the gain of electrons resulting in a

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6$ $2\text{H}_2\text{NO}_2$ $\text{C}_6\text{H}_6\text{O}_6$ $2\text{H}_2\text{O}_2$ choose 1 answer $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 a $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 b $\text{C}_6\text{H}_8\text{O}_6$ and H_2O_2 c $\text{C}_6\text{H}_8\text{O}_6$ and H_2O_2 d $\text{C}_6\text{H}_8\text{O}_6$ and H_2O_2

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