

- **Thermodynamics** is the study of energy and its transformations.
 - **First Law of Thermodynamics**: Energy cannot be created or destroyed, only transformed.

$$\Delta U = Q - W$$

 where ΔU is the change in internal energy, Q is the heat added to the system, and W is the work done by the system.
 - **Second Law of Thermodynamics**: The entropy of an isolated system always increases over time.

$$\Delta S \geq 0$$

 where ΔS is the change in entropy.



- **Enthalpy** (H) is a thermodynamic property that combines internal energy and the product of pressure and volume.

$$H = U + PV$$

 where U is internal energy, P is pressure, and V is volume.
 - **Enthalpy Change** (ΔH) is the heat transferred at constant pressure.

$$\Delta H = Q_p$$

 where Q_p is the heat added at constant pressure.



- **Free Energy** (G) is a thermodynamic potential that measures the maximum reversible work that can be performed by a system at constant temperature and pressure.

$$G = H - TS$$

 where H is enthalpy, T is temperature, and S is entropy.
 - **Free Energy Change** (ΔG) determines the spontaneity of a process.

$$\Delta G < 0$$
 (spontaneous)

$$\Delta G = 0$$
 (equilibrium)

$$\Delta G > 0$$
 (non-spontaneous)



- **Chemical Equilibrium** is a state where the concentrations of reactants and products remain constant over time.

$$K = \frac{[Products]}{[Reactants]}$$

 where K is the equilibrium constant.
 - **Le Chatelier's Principle**: If a system at equilibrium is subjected to a change in concentration, pressure, or temperature, the system will shift to counteract the change.

- **Reaction Quotient** (Q) is a measure of the relative amounts of products and reactants present during a reaction at a particular point in time.

$$Q = \frac{[Products]}{[Reactants]}$$

 where Q is the reaction quotient.

- **Free Energy Change** (ΔG) determines the spontaneity of a process.

$$\Delta G < 0$$
 (spontaneous)

$$\Delta G = 0$$
 (equilibrium)

$$\Delta G > 0$$
 (non-spontaneous)



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