# 4) Oxidation-Reduction Reactions

Oxidation - is a <u>loss</u> of electrons - is an <u>decrease</u> in oxidation number

Reduction - is a gain of electrons
- is a decreas in oxidation number

Oxidizing Agent - brings about the oxidation of another species

Reducing Agent- brings about the reduction of another species causes something else to gain electrons by losing electrons itself (Is Oxidized)

→ OIL RIG (Qxidation Is Loss; Reduction Is Gain)
LEO~GER (Lose Electrons Qxidation ~ Gain Electrons Reduction)

Memorize common oxidizing pathways (reducing goes backward)

- 1. Metal → metal ion
- Metal → metallous ion → metallic ion
- 3. Halide ion halogen halite ion halate ion (similar for No. No.)

Memorize common oxidizing / reducing agents (see bright orange sheet)

Ex. MnO₄ → Mn (acidified solution)

Ex. MnO<sub>4</sub> → Mn<sup>2+</sup> (acidified solution)

MnO<sub>4</sub> → MnO<sub>5</sub> (basic solution)

Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup> → Cr<sup>3+</sup> (acidified solution)

Cr<sub>2</sub>O<sub>7</sub><sup>2</sup> → Cr<sup>3+</sup> (acidified solution)
===Identify the substances being oxidized and reduced in each of the following ====

A) I<sub>2</sub> + IO<sub>3</sub> + H<sup>+</sup> ---> I<sup>+</sup> + H<sub>2</sub>O

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