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**Nov 27**

# Cyber Monday Last 90 Days

**ML Yell**



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negative when heat flows from the system to the surroundings thus exothermic reactions have a negative  $\Delta h$  of reaction the word exothermic is derived from exo the greek word for outside and therme the greek word

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web in the context of a chemical reaction the term neutralization is used for a reaction between an acid and a base or alkali historically this reaction was represented as acid base alkali salt water for example  $\text{hcl naoh nacl h}_2\text{o}$

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web during this reaction the  $\text{h}^+$  ions from the acid and the  $\text{oh}^-$  ions from the alkali react together to form water as this happens energy is released which is detected by the thermometer

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web other articles where exothermic reaction is discussed alkali metal reactions with nonmetals the reactions are highly exothermic producing up to 235 kcal mole for lithium fluoride the alkali metals react with nonmetals in groups 15 and 16 va and via of the periodic table sulfides can be formed by the direct reaction of the alkali metals with

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web 1 sodium hydroxide solution is poured into a beaker of hydrochloric acid which contains a thermometer showing room temperature 2 the beaker now contains sodium chloride and water and the

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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable measure an appropriate volume of each liquid eg 25 cm 3 place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kj less exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with hcl as 57 9 kj mol 1

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eg 1  $\text{hcl aq} + \text{naoh aq} \rightarrow \text{nacl aq} + \text{h}_2\text{o l}$  for which  $\Delta h = 57\text{ kJ mol}^{-1}$  eg 2  $\text{h}_2\text{so}_4 \text{ aq} + \text{koh aq} \rightarrow \text{kn o}_3 \text{ aq} + \text{h}_2\text{o l}$  for which  $\Delta h = 57\text{ kJ mol}^{-1}$   
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