

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong odor. They are prepared by the reduction of nitriles.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Epoxide Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Free-Radical Reactions

There are two types of free-radical reactions: chain and non-chain reactions.

- Chain reactions are those in which the reaction proceeds by a series of steps, each of which produces a free radical.
- Non-chain reactions are those in which the reaction proceeds by a single step.



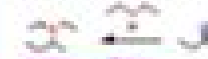
- Covalent bonds are formed by sharing of electrons between two atoms.



The Relative Strength of Covalent Bonds



Bond and Reaction Rates



Bond Energy



There are two types of free-radical reactions

1. Chain reactions



2. Non-chain reactions



Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds.

- Sigma bonds are formed by the end-to-end overlap of orbitals.
- Pi bonds are formed by the side-to-side overlap of orbitals.

- A single bond is a sigma bond.
- A double bond consists of one sigma and one pi bond.



pi bonds are weaker than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity of the atoms, covalent bonds can be polar and nonpolar.

As a rough measure,

- If the difference in electronegativity is less than 0.5, the electrons are shared equally, and the bond is nonpolar.
- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally, and the bond is polar.
- If the difference in electronegativity is greater than 1.7, the electrons are not shared, and the bond is ionic.

- A difference of 1.7 or higher can be high, but the difference can be lower, and still be a polar bond.

- A difference of 1.7 or higher can be high, but the difference can be lower, and still be a polar bond.

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.
- The more electronegative atom has a partial negative charge (δ^-), and the less electronegative atom has a partial positive charge (δ^+).
- δ^- (delta minus) indicates an electron-deficient atom.
- δ^+ (delta plus) indicates an electron-rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar, even if the dipoles are not zero.



Since the vector sum of the dipoles is zero, the molecule is nonpolar.

Bond Length and Bond Strength

The length of a bond is the distance between the nuclei of the two atoms.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds, and triple bonds are stronger than double bonds.

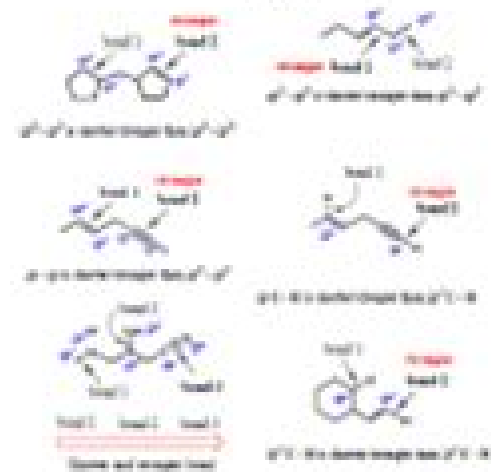


The length of a bond is the distance between the nuclei of the two atoms.

	sp ³ -sp ³	sp ² -sp ²	sp-sp
Length	154 pm	134 pm	121 pm
Strength	376 kJ/mol	611 kJ/mol	837 kJ/mol
Hybridization	sp ³ -sp ³	sp ² -sp ²	sp-sp
Electronegativity	2.1	2.5	3.5
Length	154 pm	134 pm	121 pm
Strength	376 kJ/mol	611 kJ/mol	837 kJ/mol

Higher electronegativity and stronger bond

Examples



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Brendan G. Carr



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