

# 

- **Thermodynamics** is the study of energy and its transformations.  
 - **First Law of Thermodynamics**: Energy cannot be created or destroyed, only transformed.  

$$\Delta U = Q - W$$
  
 where  $\Delta U$  is the change in internal energy,  $Q$  is the heat added to the system, and  $W$  is the work done by the system.  
 - **Second Law of Thermodynamics**: The entropy of an isolated system always increases over time.  

$$\Delta S \geq 0$$
  
 where  $\Delta S$  is the change in entropy.



- **Enthalpy** ( $H$ ) is a thermodynamic property that represents the total energy of a system.  

$$H = U + PV$$
  
 where  $U$  is internal energy,  $P$  is pressure, and  $V$  is volume.  
 - **Enthalpy Change** ( $\Delta H$ ) is the change in enthalpy during a process.  
 - **Heat Capacity** ( $C$ ) is the amount of heat required to raise the temperature of a substance by one degree.  

$$Q = C \Delta T$$
  
 where  $Q$  is the heat added,  $C$  is the heat capacity, and  $\Delta T$  is the change in temperature.



- **Free Energy** ( $G$ ) is a thermodynamic property that represents the maximum work that can be extracted from a system.  

$$G = H - TS$$
  
 where  $H$  is enthalpy,  $T$  is temperature, and  $S$  is entropy.  
 - **Free Energy Change** ( $\Delta G$ ) is the change in free energy during a process.  
 - **Equilibrium** is reached when the free energy of a system is at a minimum.



- **Chemical Equilibrium** is a state where the concentrations of reactants and products are constant.  

$$K = \frac{[Products]}{[Reactants]}$$
  
 where  $K$  is the equilibrium constant, and the brackets represent the concentrations of the species.  
 - **Le Chatelier's Principle** states that if a system at equilibrium is subjected to a change in concentration, pressure, or temperature, the system will shift to counteract the change.

- **Reaction Quotient** ( $Q$ ) is a measure of the relative amounts of reactants and products in a reaction.  

$$Q = \frac{[Products]}{[Reactants]}$$
  
 where  $Q$  is the reaction quotient, and the brackets represent the concentrations of the species.

- **Free Energy Change** ( $\Delta G$ ) is related to the equilibrium constant ( $K$ ) by the equation:  

$$\Delta G = -RT \ln K$$
  
 where  $R$  is the gas constant,  $T$  is the temperature, and  $\ln$  is the natural logarithm.



# Microeconomics Study Guide Final

**Hussin A.Rothana**



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