

KEY

1. Complete and classify the following partial equations

- a) $1 \text{ Fe} + 3 \text{H}_2\text{SO}_4 \rightarrow 1 \text{ FeSO}_4 + 3 \text{H}_2$ (single repl)
 b) $\text{NaI} + \text{Cl}_2 \rightarrow$
 c) $1 \text{ Zn} + 1 \text{ Pb(NO}_3)_2 \rightarrow 1 \text{ Zn(NO}_3)_2 + 1 \text{ Pb}$ (single repl)
 d) $\text{Al}_2\text{O}_3 + \text{HCl} \rightarrow$
 e) $2 \text{ Cu} + \text{O}_2 \rightarrow 2 \text{ CuO}$ or $4 \text{ Cu} + 1 \text{ O}_2 \rightarrow 2 \text{ Cu}_2\text{O}$ (synthesis)
 f) $\text{Al}_2\text{O}_3 \rightarrow$
 g) $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$ (synth)
 h) $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow$
 i) $1 \text{ Mg} + 2 \text{ HNO}_3 \rightarrow 1 \text{ Mg(NO}_3)_2 + 1 \text{ H}_2$ (s.r.)
 j) $\text{Na} + \text{H}_2\text{O} \rightarrow$
 k) $3 \text{ NaOH} + \text{H}_3\text{PO}_4 \rightarrow 1 \text{ Na}_3\text{PO}_4 + 3 \text{ H}_2\text{O}$ (acid base)
 l) $\text{H}_2\text{SO}_4 + \text{Al(OH)}_3 \rightarrow$
 m) $2 \text{ C}_2\text{H}_6 + 7 \text{ O}_2 \rightarrow 4 \text{ CO}_2 + 6 \text{ H}_2\text{O}$ (combustion)
 n) $\text{AlCl}_3 \rightarrow$
 o) $2 \text{ KOH} + \text{H}_2\text{SO}_4 \rightarrow 1 \text{ K}_2\text{SO}_4 + 2 \text{ H}_2\text{O}$ (acid base)
 p) $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow$
 q) $1 \text{ NaOH} + \text{HCl} \rightarrow 1 \text{ NaCl} + 1 \text{ H}_2\text{O}$ (acid base)
 r) $\text{NaHCO}_3 + \text{H}_2\text{SO}_4 \rightarrow$
 s) $1 \text{ C}_2\text{H}_6 + 7 \text{ O}_2 \rightarrow 4 \text{ CO}_2 + 6 \text{ H}_2\text{O}$ (combustion)
 t) $\text{H}_2\text{O} \rightarrow$
 u) $2 \text{ HNO}_3 + \text{CaCO}_3 \rightarrow 1 \text{ Ca(NO}_3)_2 + 1 \text{ H}_2\text{O}$ (s.r.)
 v) $\text{CH}_4 + \text{O}_2 \rightarrow$
 w) $1 \text{ Pb} + 1 \text{ H}_2\text{SO}_4 \rightarrow 1 \text{ PbSO}_4 + 1 \text{ H}_2$ (s.r.)
 x) $\text{KI} + \text{Br}_2 \rightarrow$
 y) $1 \text{ Ba(OH)}_2 + 2 \text{ HNO}_3 \rightarrow 1 \text{ Ba(NO}_3)_2 + 2 \text{ H}_2\text{O}$ (acid base)
 z) $\text{Pb} + \text{Cu(NO}_3)_2 \rightarrow$

2. Given: $2 \text{ NaClO}_3 \rightarrow 2 \text{ NaCl} + 3 \text{ O}_2$ calculate the mass of oxygen gas produced if 143 g of NaClO_3 is decomposed.

65.4g

3. Calculate the mass of oxygen gas required to burn 242 g of wax ($\text{C}_{21}\text{H}_{44}$)

254g

4. If 14.0 moles of butane (C_4H_{10}) is burned

- a) calculate the number of moles of oxygen gas required. 90.0 mol
 b) calculate the number of moles of carbon dioxide produced. 56.0 mol
 c) calculate the mass of water produced. 1080 g

5. Metallic aluminum reacts with zinc nitrate to produce metallic zinc and aluminum nitrate.

- a) Calculate the mass of aluminum that will react with 1.30 moles of zinc nitrate. 27.4g Al
 b) Calculate the mass of zinc that would be produced from 50.0 g of aluminum and sufficient zinc nitrate. 182 g Zn

6. Given: $\text{Cu} + 2 \text{ AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2 \text{ Ag}$

- a) How many grams of Cu are required to react with 3.00 g of silver nitrate? 0.544g Cu
 b) How many grams of silver are produced from 3.00 g of silver nitrate and sufficient Cu? 1.70g Ag

2) NaClO_3 143g
 $\text{NaClO}_3 \rightarrow \text{NaCl} + \frac{3}{2} \text{O}_2$
 $143 \text{ g} \times \frac{1}{106.44} = 1.344 \text{ mol}$
 $1.344 \text{ mol} \times \frac{3}{2} = 2.016 \text{ mol O}_2$
 $2.016 \text{ mol} \times 32 \text{ g/mol} = 64.512 \text{ g}$
 or
 65.4 g



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Nollet, Fidel Toldrá, Y. H. Hui, 2008-02-28 Food diagnostics is a relatively new and emerging area fuelled in large part by the ever increasing demand for food safety. *Advances in Food Diagnostics* provides the most updated comprehensive professional reference source available covering sophisticated diagnostic technology for the food industry. Editors Nollet Toldrá and Hui and their broad team of international contributors address the most recent advances in food diagnostics through multiple approaches reviewing novel technologies to evaluate fresh products describing and analyzing in depth several specific modern diagnostics providing an analysis of data processing and discussing global marketing with an insight into future trends. While covering conventional typically lab based methods of analysis the book focuses on leading edge technologies that are being or about to be introduced. The book looks at areas such as food quality assurance safety and traceability. Issues such as improved quality control monitoring pesticide and herbicide residues in food determining the nutritional content of food and distinguishing between GM and conventional foodstuffs are covered. *Advances in Food Diagnostics* offers the food professional what its title promises the latest advances in food diagnostics and analysis. [Nuclear Science Abstracts](#), 1950

Healthy Healing Linda Page, 2004-09 The Ultimate Resource For Improving Your Health Naturally Over 1 million copies sold. In its first edition nearly 20 years ago Dr Linda Page's book *Healthy Healing* was the only one of its kind. Now updated and expanded *Healthy Healing* is still the easiest to use bestselling natural health reference book on the market. Customize your own personal healing program using natural therapies for more than 300 ailments through diet whole herb supplements and exercise. Live Longer feel better and look better naturally. **A Survey of American Chemistry** National Research Council (U.S.). Division of Chemistry and Chemical Technology, 1933

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