

temperature:

- a measure of the average kinetic energy of the particles in a sample of matter
- does not depend on the amount of matter in the sample
- symbol is T; unit is °C

heat:

- total amount of kinetic energy that flows because of a difference in temperature
- depends on mass of sample
- symbol is Q; unit is J (1 J = 4.18 cal)

Kinetic energy is energy in motion

Potential energy is store

- Potential energy is hiding and cannot be measured
- Only changes in P.E. can be measured.

specific heat capacity:

- amount of energy required to raise the temperature of 1 gram of substance 1 °C
- symbol is c; unit is J/g°C

$$\begin{array}{c}
 \text{mass} \swarrow \\
 Q = m \times C \times \Delta t \\
 \swarrow \quad \quad \quad \nwarrow \\
 \text{heat} \quad \quad \quad \text{specific heat} \quad \quad \quad \text{change in temperature}
 \end{array}$$

When heat (Q) is absorbed by a system, part of it (C) goes into storage as potential energy and part of it is used to make the molecules move around more, raising the temperature (Δt).

Why does sand get hotter in the day and colder at night than the water?

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G Orfield



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Table of Contents Note Taking Guide Episode 1301 Physics Fundamentals

1. Understanding the eBook Note Taking Guide Episode 1301 Physics Fundamentals
 - The Rise of Digital Reading Note Taking Guide Episode 1301 Physics Fundamentals
 - Advantages of eBooks Over Traditional Books
2. Identifying Note Taking Guide Episode 1301 Physics Fundamentals
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Note Taking Guide Episode 1301 Physics Fundamentals
 - User-Friendly Interface
4. Exploring eBook Recommendations from Note Taking Guide Episode 1301 Physics Fundamentals
 - Personalized Recommendations
 - Note Taking Guide Episode 1301 Physics Fundamentals User Reviews and Ratings
 - Note Taking Guide Episode 1301 Physics Fundamentals and Bestseller Lists

5. Accessing Note Taking Guide Episode 1301 Physics Fundamentals Free and Paid eBooks
 - Note Taking Guide Episode 1301 Physics Fundamentals Public Domain eBooks
 - Note Taking Guide Episode 1301 Physics Fundamentals eBook Subscription Services
 - Note Taking Guide Episode 1301 Physics Fundamentals Budget-Friendly Options
6. Navigating Note Taking Guide Episode 1301 Physics Fundamentals eBook Formats
 - ePub, PDF, MOBI, and More
 - Note Taking Guide Episode 1301 Physics Fundamentals Compatibility with Devices
 - Note Taking Guide Episode 1301 Physics Fundamentals Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Note Taking Guide Episode 1301 Physics Fundamentals
 - Highlighting and Note-Taking Note Taking Guide Episode 1301 Physics Fundamentals
 - Interactive Elements Note Taking Guide Episode 1301 Physics Fundamentals
8. Staying Engaged with Note Taking Guide Episode 1301 Physics Fundamentals
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Note Taking Guide Episode 1301 Physics Fundamentals
9. Balancing eBooks and Physical Books Note Taking Guide Episode 1301 Physics Fundamentals
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Note Taking Guide Episode 1301 Physics Fundamentals
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Note Taking Guide Episode 1301 Physics Fundamentals
 - Setting Reading Goals Note Taking Guide Episode 1301 Physics Fundamentals
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Note Taking Guide Episode 1301 Physics Fundamentals
 - Fact-Checking eBook Content of Note Taking Guide Episode 1301 Physics Fundamentals
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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