

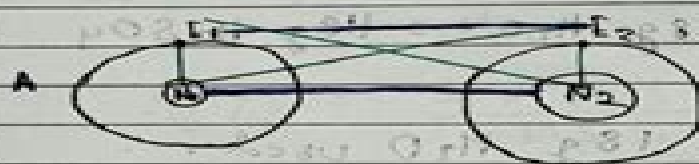
Chemical Bonding:

Chemical Bond → The attractive force which holds various constituents (atom, ion, etc) together in different chemical species is called chemical bond.

Cause of Chemical Bond:

To gain stability

- 1) Tendency to acquire minimum energy.
- 2) Tendency to acquire Noble gas configuration.



Attractive force

$N_1 + E_1$
 $N_2 + E_2$
 $N_1 + E_2$
 $N_2 + E_1$

Repulsive force

$E_1 + E_2$
 $N_1 + N_2$

* force of attraction

So, forms chemical Bond.

force of repulsion

Where attraction > repulsion

↑ attraction = energy ↓ = stability ↑

↓ attraction = energy ↑ = stability ↓

* force of attraction = force of repulsion (energy min)

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Gernot Frenking, Sason Shaik



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Chemical Bonding for JEE Main & Advanced, NEET 2nd Edition Vaibhav Trivedi, 2019-10-10 The thoroughly revised updated 2nd edition of the book on Chemical Bonding is designed especially in accordance with latest competitive trends The book has been updated with the past questions of NEET JEE Main JEE Advanced A new chapter entitled Hydrolysis of Covalent Compounds has been added based on student s high demand The salient features of the book are as follows A moderately concise and compact book covering all topics from A Z Bent Rule with latest amendments and Drago s Rule Physical properties of ionic covalent compounds with detailed explanation Increasing and decreasing order of lattice energy hydration energy polarization and effect of these on physical properties has been done comparatively Simple language to make it useful even to average and weak students Logical and evolutionary approach in descriptions for better imagination and visualization Large no of solved examples illustrations and Objective type questions Miscellaneous Practice Problems as final challenge

New Pattern Physical Chemistry For Iit-Jee Er. Shishir Mittal, This book is an attempt to present Physical Chemistry in a new innovative and an interesting manner The whole text has been divided into 40 chapters Mole Concept the heart of Physical Chemistry has been discussed in 10 chapters Important topics like Atomic Structure Chemical Bonding Gaseous State Electrochemistry Ionic Equilibrium Thermodynamics have further been divided into 2 4 chapters each so as to discuss them exhaustively The salient features of the book are i Extensive and well written theory ii Wide level of thought provoking questions and problems in the form of Concept Building Exercises Concept Testing Exercises and Objective Exercises iii The logical order or sequencing of solved examples and questions iv A new and innovative presentation supported by innumerable number of figures Data tables pictures computer generated graphs or structures analogies and Interesting related facts in boxes pictorial representation of crucial concepts This is done to make learning easy and interesting About the Author Shishir Mittal received his B Tech in chemical engineering from IIT Bombay and is currently training Students for IIT JEE in Kota

Super Course in Chemistry for the IIT-JEE: Physical Chemistry, *Chemical Bonds* Jeremy K. Burdett, 1997-05-21 Taking a question and answer format this tutorial explores the fundamental concepts of chemical bonding and is based on the lectures of Professor Jeremy Burdett

Lectures On Chemical Bonding And Quantum Chemistry Sambhu N Datta, 2019-02-25 The concept of a chemical bond evolved from a variety of experimental observations It became useful to understand at times even predict the molecular structure reactivity and mechanism of chemical reactions Every aspect of the concept of bonding received a quantitative interpretation from the advent of quantum mechanics and its application to chemistry In Lectures on Chemical Bonding and Quantum Chemistry the reader will find a comprehensive discourse on the basic interpretation of the chemical bond as well as current understanding in terms of a dancing molecule that not only travels rotates and pulsates around an equilibrium molecular structure but also interacts and collides with other molecules thereby transferring linear and angular momentum characteristics and adjusting total energies

One will also find a thorough survey of quantum mechanical methodologies for calculation of molecular characteristics in specific states and their changes under spectroscopic transitions tunneling electron and proton transfer phenomena and so on Guides to more advanced levels of theory are also provided

The Chemical Bond Gernot Frenking, Sason Shaik, 2014-07-08 This is the perfect complement to *Chemical Bonding Across the Periodic Table* by the same editors who are two of the top scientists working on this topic each with extensive experience and important connections within the community The resulting book is a unique overview of the different approaches used for describing a chemical bond including molecular orbital based valence bond based ELF AIM and density functional based methods It takes into account the many developments that have taken place in the field over the past few decades due to the rapid advances in quantum chemical models and faster computers

Chemical Bonding and the Periodic Table, 2000 *Atomic Structure and Chemical Bonding, a Non-mathematical Introduction* Fritz Seel, 1963

Atoms, Chemical Bonds, and Bond Dissociation Energies Sándor Fliszár, 1994 Chemical bonds their intrinsic energies in ground state molecules and the energies required for their actual cleavage are the subject of this book The theory modelled after a description of valence electrons in isolated atoms explains how intrinsic bond energies depend on the amount of electronic charge carried by the bond forming atoms It also explains how bond dissociation depends on these charges While this theory vividly explains thermochemical stability future research could benefit from a better understanding of bond dissociation if we learn how the environment of a molecule affects its charges we also learn how it modifies bond dissociation in that molecule This essay is aimed at theoretical and physical organic chemists who are looking for new perspectives to old problems

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