

- c. $S_2O_8^{2-}$ has SO_4 units with an angle of about 90° between their planes, in an eclipsed conformation.



- d. N_2H_4 has a trigonal pyramidal shape at each N, and a gauche conformation. There is one lone pair on each N.

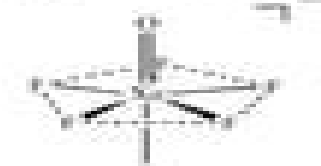


- e. ClO_2^- is a distorted trigonal pyramidal with one lone pair and double bond character in the Cl—O bond.



- f. CS_2 , like CO_2 , is linear with double bonds. (Same comment.)

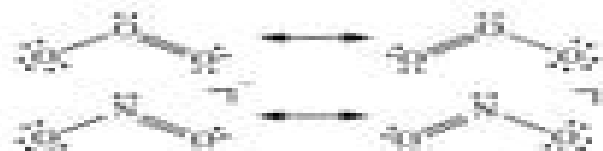
- g. The structure of NaO_5^{2-} is based on a pentagonal bipyramidal, with a lone pair and the oxygen atom in axial positions. See R. O. Christie et al., *Inorg. Chem.*, **1995**, *34*, 1555 for evidence in support of this structure.



3.11 All the halate ions are trigonal pyramids; as the central atom increases in size, the bonding pairs are farther from the center, and the lone pair forces a smaller angle. The decreasing electronegativity $Cl > Br > I$ of the central atom also allows the electrons to be pulled farther out, reducing the lps-lps repulsion.

- 3.12** a. AsH_3 should have the smallest angle, since it has the largest central atom. This minimizes the bond pair—bond pair repulsions and allows a smaller angle. Arsenic is also the least electronegative central atom, allowing the electrons to be drawn out farther and lowering the repulsions further. Actual angles: $AsH_3 = 91.8^\circ$, $PH_3 = 93.8^\circ$, $NH_3 = 106.6^\circ$.
- b. Cl is larger than F, and F is more electronegative and should pull the electrons farther from the S, so the F—S—F angle should be smaller in OSF_2 . This is consistent with the experimental data: the F—S—F angle in OSF_2 is 92.3° and the Cl—S—Cl angle in $OSCl_2$ is 96.2° .

- c. NO_2^- has rather variable angles (115° and 132°) in different salts. The sodium salt (115.4°) has a slightly smaller angle than O_3 (116.8°). The N—O electronegativity difference should pull electrons away from N, reducing the lps-lps repulsion and the angle.



- d. BrO_3^- (104°) has a slightly smaller angle than ClO_3^- (107°), since it has a larger central atom. In addition, the greater electronegativity of Cl holds the electrons closer and increases lps-lps repulsion.

3.13 a. N_2^{+} is linear, with two double bonds. O_3 is bent (see solution to 3.12 c), with one double bond and a lone pair on the central O caused by the extra pair of electrons.

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Gary D. Christion



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